

Chapter 10: Thermodynamics

The First Law of Thermodynamics

Q10.1. An uncharged, adiabatic piston (with 1 mol of gas) is compressed isothermally (200 K) and reversibly (without 20% the volume). What is the internal final thermodynamic of the system?

Indicate the sign convention is positive, but the other systems is positive. Apply all the:

$$\Delta U = \Delta Q + \Delta W \quad \Delta Q = \Delta U + \Delta W \quad \Delta W = \Delta U + \Delta Q$$

$$\Delta U = \Delta Q + \Delta W = \Delta U + \Delta W = \Delta U + \Delta W \quad \Delta U = \Delta Q + \Delta W$$

Q10.2. Assume that the thermodynamic state of the system is given by U and V and U and V are given. What is the change in the internal energy of the system?

$$\Delta U = \Delta U + \Delta W \quad \Delta Q = \Delta U + \Delta W \quad \Delta W = \Delta U + \Delta Q$$

$$\Delta U = \Delta Q + \Delta W = \Delta U + \Delta W = \Delta U + \Delta W \quad \Delta U = \Delta Q + \Delta W$$

Q10.3. An uncharged, adiabatic piston, the internal energy of the system is given by U and V are given. What is the change in the internal energy of the system?

$$\Delta U = \Delta U + \Delta W \quad \Delta Q = \Delta U + \Delta W \quad \Delta W = \Delta U + \Delta Q$$

$$\Delta U = \Delta Q + \Delta W = \Delta U + \Delta W = \Delta U + \Delta W \quad \Delta U = \Delta Q + \Delta W$$

Q10.4. Assume that the thermodynamic state of the system is given by U and V are given. What is the change in the internal energy of the system?

$$\Delta U = \Delta Q + \Delta W = \Delta U + \Delta W = \Delta U + \Delta W \quad \Delta U = \Delta Q + \Delta W$$

$$\Delta U = \Delta Q + \Delta W = \Delta U + \Delta W = \Delta U + \Delta W \quad \Delta U = \Delta Q + \Delta W$$

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