

Chapter 21 Nuclear Chemistry

Review Definitions

Chemical Energy is energy stored in the bonds of atoms and molecules.

Nuclear Energy is due to changes in the nucleus of atoms changing them into different atoms. Nuclear energy is enormous in comparison.

The Nucleus



Nucleus is composed of the two **nucleons**, protons and neutrons.

The number of protons is the atomic number.

The number of protons and neutrons together is the mass number.

Isotopes: the same element have the different mass, due to different numbers of neutrons in those atoms.

Radioactivity: some nuclides of an elements are unstable, or radioactive, which can be called **radionuclides** (放射性核素).

Table 21.1 Properties of Alpha, Beta, and Gamma Radiation			
Property	Type of Radiation		
	α	β	γ
Charge	2+	1-	0
Mass	$6.64 \times 10^{-24}\text{g}$	$9.11 \times 10^{-28}\text{g}$	0
Relative penetrating power	1	100	10,000
Nature of radiation	He nuclei	Electrons	High-energy photons

Types of Radioactive Decay:

1. **Alpha decay** is the loss of an α -particle.



2. **Beta decay** is the loss of a β -particle (a high-speed electron emitted by the nucleus): $^0_{-1}\beta$ or $^0_{-1}\text{e}$



3. **Gamma emission** is the loss of a γ -ray, which is high-energy radiation that almost always accompanies the loss of a nuclear particle: $^0_0\gamma$

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Table of Contents Chapter 21 Nuclear Chemistry Worksheets

1. Understanding the eBook Chapter 21 Nuclear Chemistry Worksheets
 - The Rise of Digital Reading Chapter 21 Nuclear Chemistry Worksheets
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 21 Nuclear Chemistry Worksheets
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 21 Nuclear Chemistry Worksheets
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 21 Nuclear Chemistry Worksheets
 - Personalized Recommendations
 - Chapter 21 Nuclear Chemistry Worksheets User Reviews and Ratings

- Chapter 21 Nuclear Chemistry Worksheets and Bestseller Lists
- 5. Accessing Chapter 21 Nuclear Chemistry Worksheets Free and Paid eBooks
 - Chapter 21 Nuclear Chemistry Worksheets Public Domain eBooks
 - Chapter 21 Nuclear Chemistry Worksheets eBook Subscription Services
 - Chapter 21 Nuclear Chemistry Worksheets Budget-Friendly Options
- 6. Navigating Chapter 21 Nuclear Chemistry Worksheets eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 21 Nuclear Chemistry Worksheets Compatibility with Devices
 - Chapter 21 Nuclear Chemistry Worksheets Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 21 Nuclear Chemistry Worksheets
 - Highlighting and Note-Taking Chapter 21 Nuclear Chemistry Worksheets
 - Interactive Elements Chapter 21 Nuclear Chemistry Worksheets
- 8. Staying Engaged with Chapter 21 Nuclear Chemistry Worksheets
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 21 Nuclear Chemistry Worksheets
- 9. Balancing eBooks and Physical Books Chapter 21 Nuclear Chemistry Worksheets
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 21 Nuclear Chemistry Worksheets
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 21 Nuclear Chemistry Worksheets
 - Setting Reading Goals Chapter 21 Nuclear Chemistry Worksheets
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 21 Nuclear Chemistry Worksheets
 - Fact-Checking eBook Content of Chapter 21 Nuclear Chemistry Worksheets
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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