

CHAPTER 25

NUCLEAR CHEMISTRY

PRACTICE EXAMPLES

1A (E) A β^- has a mass number of zero and an “atomic number” of -1 . Emission of this electron has the effect of transforming a neutron into a proton. ${}_{94}^{241}\text{Pu} \rightarrow {}_{95}^{241}\text{Am} + {}_{-1}^0\beta$

1B (E) ${}^{58}_{28}\text{Ni}$ has a mass number of 58 and an atomic number of 28. A positron has a mass number of 0 and an effective atomic number of $+1$. Emission of a positron has the seeming effect of transforming a proton into a neutron. The parent nuclide must be copper-58.



2A (E) The sum of the mass numbers ($139 + 12 = ? + 147$) tells us that the other product species has $A = 4$. The atomic number of La is 57, that of C is 6, and that of Eu is 63. The atomic number sum ($57 + 6 = ? + 63$) indicates that the atomic number of this product species is zero. Therefore, four neutrons must have been emitted.



2B (E) An alpha particle is ${}_2^4\text{He}$ and a positron is ${}_{+1}^0\beta$. We note that the total mass number in the first equation is 125; the mass number of the additional product is 1. The total atomic number is 53; the atomic number of the additional product is 0; it is a neutron. ${}_{53}^{125}\text{Sb} + {}_2^4\text{He} \rightarrow {}_{54}^{129}\text{I} + {}_0^1n$

In the second equation, the positron has a mass number of 0, meaning that the mass number of the product is 124. Because the atomic number of the positron is $+1$, that of the product is 52; it is ${}_{52}^{124}\text{Te}$.



3A (M) (a) The decay constant is found from the 8.040-day half-life.

$$\lambda = \frac{0.693}{8.040 \text{ d}} = 0.0862 \text{ d}^{-1} \times \frac{1 \text{ d}}{24 \text{ h}} \times \frac{1 \text{ h}}{60 \text{ min}} \times \frac{1 \text{ min}}{60 \text{ s}} = 9.98 \times 10^{-7} \text{ s}^{-1}$$

(b) The number of ${}^{131}\text{I}$ atoms is used to find the activity.

$$\begin{aligned} \text{no. } {}^{131}\text{I} \text{ atoms} &= 2.05 \text{ mg} \times \frac{1 \text{ g}}{1000 \text{ mg}} \times \frac{1 \text{ mol } {}^{131}\text{I}}{131 \text{ g } {}^{131}\text{I}} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mol } {}^{131}\text{I}} \\ &= 9.42 \times 10^{18} \text{ atoms } {}^{131}\text{I} \end{aligned}$$

$$\text{activity} = \lambda N = 9.98 \times 10^{-7} \text{ s}^{-1} \times 9.42 \times 10^{18} \text{ atoms} = 9.40 \times 10^{12} \text{ disintegrations / second}$$

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Rachel Sandford



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web oct 6 2022 the limit state method is a method of design of reinforced concrete structures and steel structures this is one of the most popular methods of designing structures and concepts that are important for the gate 2023 exam

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