

Concentration and Molarity (M) Table

Date: _____

Non-dissociating solute

Use the concentration and molarity information to determine the missing information in the table.

Volume (mL)	Mass (g)	Molarity (M)	Volume (mL)	Mass (g)	Molarity (M)
100	20	0.200	50.0	20	0.400
200	20	0.100	100	20.0	0.200
1.00	20	0.200	0.10	2.0	0.200
50	50	1.00	50.0	5.0	0.100

Concentration and Molarity (M) Table

1. **Calculate the molarity of a solution prepared by dissolving 20.0 g of NaCl in enough water to make 100.0 mL of solution.**

2. **Calculate the mass of NaCl in 250.0 mL of a 0.100 M NaCl solution.**

3. **What volume of a 0.500 M NaCl solution contains 25.0 g of NaCl?**

4. **Calculate the molarity of a solution prepared by dissolving 10.0 g of NaCl in enough water to make 250.0 mL of solution.**

5. **Calculate the mass of NaCl in 500.0 mL of a 0.200 M NaCl solution.**

6. **What volume of a 0.100 M NaCl solution contains 25.0 g of NaCl?**

7. **Calculate the molarity of a solution prepared by dissolving 10.0 g of NaCl in enough water to make 250.0 mL of solution.**

$$0.200 \text{ mol/L} = 0.271 \text{ M} \quad 0.400 \text{ M is calculated}$$

8. **Calculate the mass of NaCl in 250.0 mL of a 0.100 M NaCl solution.**

$$0.200 \text{ mol/L} \times 0.250 \text{ L} = 0.050 \text{ mol} \quad 0.200 \text{ M is calculated}$$

9. **What volume of a 0.500 M NaCl solution contains 25.0 g of NaCl?**

$$0.200 \text{ mol/L} = 0.271 \text{ M}$$

10. **Calculate the molarity of a solution prepared by dissolving 10.0 g of NaCl in enough water to make 250.0 mL of solution.**

$$0.200 \text{ mol/L} = 0.271 \text{ M}$$

11. **Calculate the mass of NaCl in 500.0 mL of a 0.200 M NaCl solution.**

$$0.40 \text{ mol} \times 0.200 \text{ mol/L} = 0.080 \text{ mol}$$

12. **What volume of a 0.100 M NaCl solution contains 25.0 g of NaCl?**

$$0.40 \text{ mol} \times 0.200 \text{ mol/L} = 0.080 \text{ mol}$$

Concept Review Concentration And Molarity Answer Key

Martin Stuart Silberberg



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