NUCLEAR CHEMISTRY - WORKSHEET

 When stable, nonradioactive nitrogen in the air is subject to neutron bombardment during a hydrogen bomb explosion, a single neutron is absorbed by the nitrogen nucleus and a proton is emitted (given off). Write a balanced nuclear equation for this reaction and name the other product of the reaction.

A radioactive element. A, with a mass of 212 and atomic number 84, emits an alpha. particle and changes into element B. Element B emits a beta particle and is converted to element C.

a) Write nuclear equations for the formation of elements B and C.

208Pb → _0e + 208Bi (note that this process is not fast; lead is stable.)

b) What are the atomic masses and atomic number of elements A. B and C?

Masses are on top: atomic numbers at bottom.

c) Name elements A. B and C.

Polonium, Lead, and Bismuth.

Complete the following nuclear equations and state the general type:

c)
$$^{120}_{49}In \rightarrow ^{0}_{-1}e + ?$$

e)
$${}^{12}C + ? \rightarrow {}^{13}N + {}^{0}V$$

b)
$$\frac{214}{51}Bi \rightarrow \frac{4}{5}He + ?$$

d)
$$_{69}^{162}Tm \rightarrow _{+1}^{0}e + ?$$

f)
$${}_{19}^{31}P + {}_{2}^{4}He \rightarrow {}_{12}^{34}CI + ?$$

a) _1β or _1e	b) 210Tl	
c) 120Sn	d) 162Er	
e) !H	f) ½n	

Write a balanced nuclear equation for each of the following:

a)
$$\frac{216}{85}At(d,\alpha)$$

a)
$$^{216}_{85}$$
At $\rightarrow ^{4}_{2}$ He + $^{212}_{83}$ Bi
b) $^{208}_{81}$ Tl $\rightarrow ^{0}_{-1}$ e + $^{282}_{88}$ Pb

b)
$$^{208}_{81}\text{Tl} \rightarrow ^{0}_{-1}\text{e} + ^{208}_{82}\text{Pb}$$

b)
$$^{1}_{81}\text{FI} \rightarrow ^{1}_{-1}\text{e} + ^{1}_{82}\text{Pb}$$

c) $^{55}_{25}\text{Mn} + ^{1}_{0}\text{n} \rightarrow ^{56}_{25}\text{Mn} + \gamma$

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