#### Solutions and Solubility

#### Dissolving Process

- Solvent molecules are attracted to solute molecules or ions at the surface of the particles.
- Solvent molecules surround the solute molecule.
- 3. Solvent molecules carry the solute molecule into the solution.
- 4. The solute molecule dissolves in a "like" solvent.

#### Factors that Affect the Rate of Dissolution

- The rate of dissolution describes the rate or speed at which a solid dissolves a liquid.
- It is affected by:
  - o Stirring
  - Surface area.
  - Temperature.

#### Solubility:

- Solubility is the amount of a solute that will dissolve in a given volume of solvent at a
  given temperature and pressure.
- It depends on the solute and the solvent.
  - Solubility does not equal dissolution rate.

#### Unsaturated, Saturated, and Supersaturated Solutions

- Types of solutions:
  - Unsaturated solution the concentration of solute is less than the maximum concentration predicted from the solute's solubility.
  - Saturated solution the concentration of solute is equal to the maximum concentration predicted from the solute's solubility, the solute crystallizes at the same rate that it dissolves.
  - Supersaturated solution the concentration of solute is greater than the maximum concentration predicted from the solute's solubility.

#### The Effect of Temperature on Solubility

- Increasing temperature:
  - Increases the solubility of solids and liquids.
  - Decreases the solubility of gases.

#### The Effect of Pressure on Solubility

- Increasing pressure:
  - Increases the solubility of gases.
  - Has little effect on the solubility of solids and liquids.

#### Solubility Graphs

- Solubility graphs:
  - Show how solubility changes with temperature.
  - Are determined experimentally.

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Lawrence K. Wang, Chih Ted Yang

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