

3. a) The mean kinetic energy of a nitrogen molecule is the same as the mean kinetic energy of an oxygen molecule since  $E_k$  is directly proportional to temperature  $E_k = \frac{3}{2} kT$

b) molar mass  $N_2 = 0.028 \text{ kg mol}^{-1}$  molar mass  $O_2 = 0.032 \text{ kg mol}^{-1}$   
 $N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$

$$\frac{1}{2} M_N C_{\text{rms},N}^2 = \frac{1}{2} M_O C_{\text{rms},O}^2$$

$$\frac{0.028}{6.02 \times 10^{23}} C_{\text{rms},N}^2 = \frac{0.032}{6.02 \times 10^{23}} C_{\text{rms},O}^2$$

$$0.028 C_{\text{rms},N}^2 = 0.032 C_{\text{rms},O}^2$$

$$C_{\text{rms},N}^2 = \frac{0.032}{0.028} C_{\text{rms},O}^2$$

$$C_{\text{rms},N}^2 = 1.14 C_{\text{rms},O}^2$$

$$C_{\text{rms},N} = \sqrt{1.14} C_{\text{rms},O}$$

$$C_{\text{rms},N} = 1.07 C_{\text{rms},O}$$

4. molar mass =  $0.028 \text{ kg mol}^{-1}$   $V = 0.037 \text{ m}^3$   $T = 300 \text{ K}$   
 $p = 100 \text{ kPa} = 100 \times 10^3 \text{ Pa}$

a)  $pV = nRT \Rightarrow \frac{pV}{RT} = n = \frac{(100 \times 10^3)(0.037)}{8.31(300)}$   
 $= \frac{3700}{2493} = 1.484 = \underline{1.48 \text{ moles to 3sf}}$

b) mass = number of moles  $\times$  molar mass  
 $= 1.484 \times 0.028 = 0.041512 \text{ kg}$   
 $= \underline{4.2 \times 10^{-2} \text{ kg to 2sf}}$

(If 1.48 moles used then mass =  $0.04144 \text{ kg} = 4.1 \times 10^{-2} \text{ kg to 2sf}$ )

c) mass of one molecule  $m = \frac{\text{molar mass}}{N_A} = \frac{0.028}{6.02 \times 10^{23}} = 4.65 \times 10^{-26} \text{ kg}$

$$\frac{1}{2} M C_{\text{rms}}^2 = \frac{3}{2} kT$$

$$4.65 \times 10^{-26} C_{\text{rms}}^2 = 3 \left( 1.38 \times 10^{-23} \right) 300 = 1.242 \times 10^{-20}$$

$$C_{\text{rms}} = \frac{1.242 \times 10^{-20}}{4.65 \times 10^{-26}} = 267 \times 10^3 \text{ m s}^{-1}$$

$$C_{\text{rms}} = \sqrt{267 \times 10^3}$$

$$= 516 \text{ m s}^{-1}$$

$$= \underline{5.2 \times 10^2 \text{ m s}^{-1} \text{ to 2sf}}$$

w.3 could use  
 $pV = \frac{1}{2} N m C_{\text{rms}}^2$   
 Since  $Nm = \text{mass}$   
 and would also be  
 equal to 1.48(0.028)

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