

## UNIT 9

# COORDINATION COMPOUNDS

### 1. Introduction

Complex compounds or coordination compounds are those molecular compounds which retain their identity in solid as well as in solution are known as complex compounds. In these compounds metal atoms are bound to a number of anions.

**Example,**

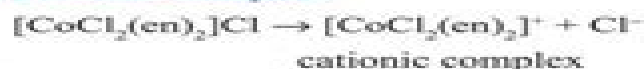


### 2. Types of Complex

#### (i) Anionic complex



#### (ii) Cationic complex



#### (iii) Neutral complex



### 3. Ligands

The ions or molecules bound to the central atom/ion in the coordination entity are called ligands.

#### Types of Ligands

- (i) Unidentate, a ligand which is bound to a metal ion through a single donor atom.  
*e.g.*,  $H_2O$ ,  $NH_3$ ,  $CO$ ,  $Cl^-$ ,  $NH_2^-$  etc.
- (ii) Didentate, a ligand which is bound to a metal ion through a two donor atoms.



# Coordination Compounds Study Guide

**Steven S. Zumdahl, Paul B. Kelter**



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