Stoichiometry Lab: Iron and Copper Sulfate Sample data

Data and observations :

mass of copper (II) sulfate pentahydrate	3.27	9
mass of iron:	0.50	9
mass of filter paper	1.28	9
mass of filter paper and dry copper:	1.84	9
mass of dry copper	0.56	9

Observations of reaction:

When the iron was added a precipitate began to form. The precipitate was reddish in color. The solution remained a blue color but was somewhat lighter in color.

Calculations:

1&2.The single replacement equation will be: Fe + CuSO₄ → FeSO₄ + Cu

Fe and CuSO₄ are the reactants and FeSO₄ and Cu are the products.

4. The moles of copper (II) sulfate pentahydrate used were 0.013 mol.

3.27 g CuSO₄ •
$$5H_2O \times \frac{1 \text{ mol CuSO}_4 \cdot 5H_2O}{249.47 \text{ g CuSO}_4 \cdot 5H_2O} = 0.013 \text{ mol CuSO}_4 \cdot 5H_2O$$

5. The moles of iron used were 0.009 mol Fe.

$$0.50 \text{ g Fe} \times \frac{1 \text{ mol Fe}}{55.85 \text{ g Fe}} = 0.009 \text{ mol Fe}$$

The moles of copper produced were 0.009 mol Cu.

$$0.56 \text{ g Cu} \times \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} = 0.009 \text{ mol Cu}$$

Copper Sulfate And Iron Lab Answers

Hussin A.Rothana

Copper Sulfate And Iron Lab Answers:

Embracing the Melody of Term: An Emotional Symphony within Copper Sulfate And Iron Lab Answers

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