

## Equilibrium Quiz

Choose the answer that best fits each question.

- Once a chemical equilibrium is established the \_\_\_\_\_ and \_\_\_\_\_ of reaction are equal to each other.  
a. initial and final concentrations      c. reactant and product masses  
b. forward and reverse rates      d. decomposition and formation processes
- Once chemical equilibrium of a reaction has been reached there is no longer a(n) \_\_\_\_\_ change in concentrations of chemical species.  
a. Thermal      b. small      c. insignificant      d. observable
- Which of the following do not affect equilibrium? Mark all that apply  
a. Gas      b. Liquid      c. Aqueous      d. Solid
- Which of the following cause an equilibrium shift? Mark all that apply  
a. Pressure      b. Concentration      c. Temperature      d. Catalyst
- Match the favorability of the reaction when  $K$  is compared to 1.  
a.  $K > 1$       i. reactants favored  
b.  $K < 1$       ii. products favored  
c.  $K = 1$       iii. equilibrium reached
- Using the reaction:  $A + B \rightleftharpoons 2C$ . If you have 4M A and 3M B then how much C will be present once the reaction reaches equilibrium?  $K = 10$   
a. 2.06 M      b. 7.36 M      c. 0.03 M      d. 4.16 M
- Using the reaction:  $A + 2B \rightleftharpoons 2C$ . If you have 2M A and 2M B then how much C will be present once the reaction reaches equilibrium?  $K = 1.0 \times 10^{-6}$   
a. 1.004 M      b. 0.0026 M      c. 0.005 M      d. 0.00023 M
- Using the reaction:  $3A + 2B \rightleftharpoons 3C$ . Calculate the equilibrium constant if the concentrations at equilibrium are  $A = 2M$ ,  $B = 3M$ , and  $C = 1.5M$   
a. 0.047      b. 0.017      c. 1.60      d. 12.01

Use the graph below to answer the questions.

# Chapter 1chemical Equilibrium Quiz

**Louis Meites**



## **Chapter 1 Chemical Equilibrium Quiz:**

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