

**RADIOACTIVITY:** spontaneous disintegration of the nucleus of certain atoms accompanied by the emission (release) of particles and/or energy

- ~ **TRANSMUTATION:** the change of one element into another due to changes in the nucleus
  - \* can be caused by radioactive decay (called \_\_\_\_\_ transmutation)
  - Particles that are emitted during radioactive decay
    - ~ alpha particle ( $\alpha$ ) - written as  ${}^4_2\text{He}$ ; low energy particle; blocked by paper, skin, clothing



- ~ beta emission ( $\beta$ ) - written as  ${}^{0}_{-1}\text{e}$  or  ${}^{0}_{-1}\beta$ ; blocked by plastic, wood, light metals



- ~ positron emission ( ${}^{0}_{+1}\beta$ ); also written as  ${}^{0}_{+1}\text{e}$



- ~ gamma radiation ( $\gamma$ ) - no mass or charge, just energy; blocked by lead or thick (1+ m) concrete

\* can be caused by bombardment by subatomic particles (called \_\_\_\_\_ transmutation)

- particle accelerator:

- ~ proton (a hydrogen nucleus)



- ~ alpha particle (a helium nucleus)



- ~ only charged particles can be accelerated - neutrons cannot be accelerated in ordinary particle accelerator

### BALANCING NUCLEAR EQUATIONS NOTES

- ~ Nuclear equations are different than regular equations because the atomic number and mass number are given along with the symbol for the element.
- ~ To balance a nuclear equation, the sum of the mass numbers of the substances on one side of the equation must equal the sum of the mass numbers of the substances on the other side of the equation. The same is true for the atomic numbers.

~ **EXAMPLE:**  ${}^{242}_{96}\text{Cm} + \underline{\hspace{1cm}} \rightarrow {}^{245}_{98}\text{Cf} + {}^1_0\text{n}$

- ~ Top # = mass number... Left side = 242, Right side = 246

So, in blank space, top number must be 4

- ~ Bottom # = atomic number... Left side = 96, Right side = 98

So, in blank space, bottom number must be 2

- ~ So far the equation is  ${}^{242}_{96}\text{Cm} + {}^4_2\underline{\hspace{0.5cm}} \rightarrow {}^{245}_{98}\text{Cf} + {}^1_0\text{n}$

- ~ Determine the element using the atomic number. He has an atomic number of 2.

- ~ So, the final equation is  ${}^{242}_{96}\text{Cm} + {}^4_2\text{He} \rightarrow {}^{245}_{98}\text{Cf} + {}^1_0\text{n}$

# Chapter 1 nuclear Chemistry Answer Key

**Peter Atkins, Loretta Jones**



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**Study Guide for Chemical Principles** Thomas Elliott Taylor, 1979

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*Comprehensive NDA/ NA Guide for Mathematics, English & General Knowledge* Disha Experts, 2020-01-24

**Prentice Hall Chemistry**, 2000

Chemistry McGraw-Hill Staff, 2001-03

**Chemistry** Oxtoby, 1994

*How to Prepare for the AP Chemistry* Neil D. Jespersen, Neil Jespersen, PH. D., 1999-02 This manual offers two diagnostic exams to help students pinpoint their strengths and weaknesses plus three full length practice exams closely modeled on actual AP chemistry exams

*Basic Concepts of Chemistry* Leo J. Malone, Theodore O. Dolter, 2012-04-13 Basic Concepts of Chemistry was originally written over thirty years ago to address the needs of general chemistry students with little or no background in chemistry Over time the text has evolved beyond purposes solely aimed at a preparatory chemistry course For some preparatory chemistry students a main

sequence in general chemistry may follow but for others a semester of organic and biochemistry may follow Other students enroll to simply satisfy a basic science or chemistry requirement The text was written with a level and functionality designed to accommodate the needs of each of these varied groups of students Basic Concepts of Chemistry was designed with a flexibility that allows instructors to emphasize or omit certain clearly delineated sections The mission of the text has evolved in response to the increased diversity of students and the emphasis on outcomes assessment

**Holt Chemistry** Ralph Thomas Myers,2004

**Arco Preparation for the GED** Seymour Barasch,1996 Includes model tests with explanations and study tips for GED

**MCAT, Medical College Admission Test** Lawrence Solomon,Morris Bramson,1988 Including more material and better coverage than any of the competition this completely revised and enlarged guide incorporates four full length sample tests complete with coverage of the new essay question

*Publications of the National Bureau of Standards ... Catalog* United States. National Bureau of Standards,1969

**McGraw-Hill's GED** McGraw-Hill Education,2001-09-11

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**VETs, Complete Preparation for the Veterinary Entrance Tests** Aftab S. Hassan,1998

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