Ch 14.6-14.7- Buffers & titrations and

Ch 15 - Equilibria of Other Reaction Classes

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Section 14.6 - Buffers

What is a buffer? What two components are required to make a buffer?

Consider a buffer that contains 0.250 M HF and 0.200 M F*.

What component in the buffer neutralizes any added base (OH1)?

What component in the buffer neutralizes any added acid (H₂O*)?

For buffer solutions, an ICE table may be used to determine equilibrium concentrations. Follow Example 14.20 (part a) to see how this is done. Create an ICE table for a buffer that is 0.250 M HF and 0.200 M F⁻.

	HF (aq)	+ H ₂ O (/)	F- (aq)	+ H ₃ O* (aq)
To .		-		
C		- I		
€.		2		

Q1: Use your ICE table to calculate the pH of this buffered solution. The K_a of HF is 3.5 \times 10⁻⁴. You may assume x is negligibly small.

The pH of **buffered solutions** can also be calculated using the Henderson-Hasselbalch equation. Write this equation below.

pH =

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13 most important help desk kpis to track and measure help desk performance - Jan 07 2023 web set databox scorecards to receive updates on key performance indicators through email mobile push notification and or slack 13 most important help desk kpis in this section we re sharing the top 13 help desk kpis that service professionals use average resolution time first response time first contact resolution rate overall customer