Oxidation - Reduction Reactions

Redox reaction: A chemical reaction in which one substance is oxidized, and another substance is reduced.

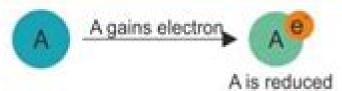
Oxidation: when an atom or group of atoms lose electron/electrons, its oxidation is said to take place.
Oxidising agent: The atom or the group of atoms which gain electron/electrons.



Example:

The above reaction is given as

Oxidation of iron: Fe loses 2 electrons Oxidising agent: S gains 2 electrons Reduction: when an atom or group of atoms gain electron/electrons, its reduction is said to take place. Reducing agent: The atom or the group of atoms which lose electron/electrons.



Example:

Reduction of Chlorine: Cl gains 2 electrons

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Robert R. Redfield

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