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## Chapter 21 Study Guide

### **Key Terms**

*Directions: Define each of the following key terms from chapters seventeen and nineteen on a separate sheet of paper.*

- **Balanced Chemical Equation** – chemical equation with the same number of atoms of each element on either side of the equation.
- **Catalyst** - speeds up a chemical reaction without being permanently changed itself.
- **Chemical Change** – another name for a chemical reaction
- **Chemical Equation** - uses chemical symbols and formulas to describe a chemical reaction and the product(s) it produces.
- **Chemical Reaction** - occur when one or more substances are changed or converted into new substances.
- **Coefficient** - numbers which represent the number of units of each substance in a reaction.
- **Combustion Reaction** - occurs when a substance reacts with oxygen to produce energy in the form of heat or light.
- **Conservation of Mass** - states matter is not created or destroyed during a chemical reaction.
- **Decomposition Reaction** - A complex substance breaks down into two or more simpler substances.
- **Double Replacement Reaction** - Different atoms in two different compounds replace each other.
- **Endergonic Reaction** – chemical reaction that requires energy input in order to proceed.
- **Endothermic Reaction** - a chemical reaction in which heat energy is absorbed.
- **Exergonic Reaction** – chemical reaction that releases some form of energy.
- **Exothermic Reaction** - a chemical reaction in which heat energy is released.
- **Inhibitor** - prevents or slows a chemical reaction or interferes with a catalyst's action.
- **Oxidation Reaction** - reaction in which electrons are lost.
- **Precipitate** – solid substance which forms as the result of a chemical reaction.
- **Product** - are the substance(s) produced by a chemical reaction. (i.e. what you end up with)
- **Reactant** - are the substance(s) that enter into a chemical reaction. (i.e. what you start with)
- **Reduction Reaction** - reaction in which electrons are gained.
- **Single Replacement Reaction** - An uncombined element replaces an element that is part of a compound.
- **Subscript** - numbers which represent the number of atoms in a molecule of a particular element.
- **Synthesis Reaction** - Two or more simple substances combine to form a new, more complex substance

# Chapter 21 Study Guide

**Anthony S. Fauci**



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