

## Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CNS: 7a)

9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ( $q = m \cdot c \cdot \Delta T$ ) & latent heat of phase change ( $q = m \cdot \Delta H$ ). (CNS: 2d)

9.5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8/\*, &)

9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

- Condensation
- Freezing
- Vaporization
- Melting
- Boiling
- Sublimation

- Deposition
- Heat
- Temperature
- Thermochemistry
  - Exothermic
  - Endothermic

- Calorimeter
- Joule
- Specific heat
- $q = m \cdot c \cdot \Delta T$
- $q = m \cdot \Delta H_{\text{vaporization}}$
- $q = m \cdot \Delta H_{\text{fusion}}$

- Heat of reaction
- Le Chatelier's principle
- Equilibrium
- Activation Energy
- Catalyst

1. According to the collision theory, what 3 things must happen in order for two molecules to react?

**Must collide with each other. Must collide with enough energy. & in the correct orientation**

2. What is the importance of the activation energy in a chemical reaction?

Amount of energy needed for a reaction to occur.

3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction:



a. Increase temp: Increase molecules collide more

b. Increase  $HCl$ : Increase more collisions because more reactants to collide

c. Increase pressure: Hard to say, wouldn't affect the reactants directly because none of them are gases.

One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call... (this one won't be on the test).

d. Remove  $MgCl_2$ : Increase, less products means less chance for them to interfere with reactant collisions

e. Add  $H_2$  : Decrease, more products means less chance for them to interfere with each other, so they would collide less.

4. Identify the following terms with their correct letter from the graphic:

• **Reactions:**

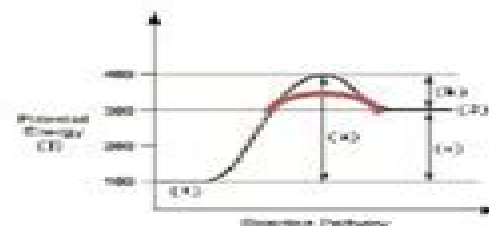
[Products](#)

- Activation energy- 

• Is this reaction exothermic or endothermic?

**Electrothermoelectric**

5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph)



# Chapter 1 reaction Rates And Equilibrium Answer Key

**John T. Moore, Richard H. Langley**



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10. *International Congress on Metallic Corrosion Part 1* Erian A. Armanios, 1988-01-01 Proceedings of the 10th International Conference on Metallic Corrosion Madras November 1987

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## **Table of Contents Chapter 1reaction Rates And Equilibrium Answer Key**

1. Understanding the eBook Chapter 1reaction Rates And Equilibrium Answer Key
  - The Rise of Digital Reading Chapter 1reaction Rates And Equilibrium Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1reaction Rates And Equilibrium Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 1reaction Rates And Equilibrium Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1reaction Rates And Equilibrium Answer Key

- Personalized Recommendations
- Chapter 1 reaction Rates And Equilibrium Answer Key User Reviews and Ratings
- Chapter 1 reaction Rates And Equilibrium Answer Key and Bestseller Lists
- 5. Accessing Chapter 1 reaction Rates And Equilibrium Answer Key Free and Paid eBooks
  - Chapter 1 reaction Rates And Equilibrium Answer Key Public Domain eBooks
  - Chapter 1 reaction Rates And Equilibrium Answer Key eBook Subscription Services
  - Chapter 1 reaction Rates And Equilibrium Answer Key Budget-Friendly Options
- 6. Navigating Chapter 1 reaction Rates And Equilibrium Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 1 reaction Rates And Equilibrium Answer Key Compatibility with Devices
  - Chapter 1 reaction Rates And Equilibrium Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 1 reaction Rates And Equilibrium Answer Key
  - Highlighting and Note-Taking Chapter 1 reaction Rates And Equilibrium Answer Key
  - Interactive Elements Chapter 1 reaction Rates And Equilibrium Answer Key
- 8. Staying Engaged with Chapter 1 reaction Rates And Equilibrium Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 1 reaction Rates And Equilibrium Answer Key
- 9. Balancing eBooks and Physical Books Chapter 1 reaction Rates And Equilibrium Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 1 reaction Rates And Equilibrium Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 1 reaction Rates And Equilibrium Answer Key
  - Setting Reading Goals Chapter 1 reaction Rates And Equilibrium Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 1 reaction Rates And Equilibrium Answer Key



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  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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




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