

Exam #10 Review: Acids, Bases, and pH

1. Know the properties of acids and bases. What are their similarities and differences?

Acids

- Taste *sour*
- React with metals to form hydrogen gas
- Contain a hydrogen ion (H^{+1})
- pH value less than 7

Bases

- Taste *bitter*
- Feel slippery
- Contain a hydroxide ion (OH^{-1})
- pH value greater than 7

BOTH

- change colors of indicators
- react with each other to form salt and water
- conduct electricity when dissolved in solution (electrolytes)

2. Know the rules for naming
- acids**
- (binary acid and oxyacid) and name the following below:

Binary acids (only 2 elements): hydro – element – ic acidOxyacids (contain a polyatomic ion): NO "HYDRO" IN THE NAME

- -ate ion → -ic acid
- -ite ion → -ous acid

REMEMBER:

"I ate something icky,
all nite I was nauseous"

- a. HCl **hydrochloric acid**
b. HF **hydrofluoric acid**
c. $HClO_2$ **chlorous acid**

- d. HNO_3 **nitric acid**
e. H_3PO_4 **phosphoric acid**
f. H_2SO_3 **sulfurous acid**

3. Know the rules for naming
- bases**
- and name the following below:

Element + hydroxide

- a. NaOH **sodium hydroxide**
b. LiOH **lithium hydroxide**

- c. $Al(OH)_3$ **aluminum hydroxide**
d. $Ba(OH)_2$ **barium hydroxide**

4. How are Arrhenius acids and bases defined?

Acids – produce H^{+1} ions in aqueous solutionsBases – produce OH^{-1} ions in aqueous solutions

5. How are Bronsted-Lowry acids and bases defined?

Acids – donate (give away) a H^{+1} ion (proton)Bases – accept (take in) a H^{+1} ion (proton)

6. Show how the following acids and bases ionize:



7. If it is a strong acid or base, how much do they ionize in solution?

STRONG acids & bases COMPLETELY ionize (all break apart) in a solution

8. If it is a weak acid or base, how much do they ionize in solution?

WEAK acids & bases do NOT completely ionize

9. What is a neutralization reaction? What products are always produced?

Acid reacts with a basealways produces water and a salt

10. Complete the following neutralization reaction and what type of salt is made:

Type of Salt: neutralType of Salt: neutral

Chapter 1review Acids Bases 1

Scott C. Dulebohn



Chapter 1review Acids Bases 1:

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