

# Stoichiometry

Solve the following problems:

1. Four moles of iron (II) sulfide added to seven moles of oxygen, produced two moles of iron (III) oxide and four moles of sulfur dioxide. (20pts)

a) Translate the problem to chemical equation.



**Reactant side**

4 Fe atoms

4 S atoms

7(2) = 14 O atoms

**Product side**

4 Fe atoms

4 S atoms

2(3) + 4(2) = 14 O atoms

b) How many grams of oxygen will react with 2.0 grams of iron (II) sulfide?

Given:

m FeS = 2.0 g

Req'd. = m O<sub>2</sub>

Solution:

molar mass of FeS = 88 g/mol

mole ratio = 4 mol FeS: 7 mol O<sub>2</sub>

molar mass of O<sub>2</sub> = 32 g/mol

$$m_{O_2} = 2.0 \cancel{g_{FeS}} \left( \frac{1 \cancel{mol_{FeS}}}{88 \cancel{g_{FeS}}} \right) \left( \frac{7 \cancel{mol_{O_2}}}{4 \cancel{mol_{FeS}}} \right) \left( \frac{32 \cancel{g_{O_2}}}{1 \cancel{mol_{O_2}}} \right) = 1.3 g_{O_2}$$

2. Rusting involves the reaction of iron, Fe with oxygen, O<sub>2</sub> to form iron (III) oxide, Fe<sub>2</sub>O<sub>3</sub>. (20pts)

a) Write the balanced chemical equation.



**Reactant side**

4 Fe atoms

3(2) = 6 O atoms

**Product side**

4 Fe atoms

2(3) = 6 O atoms

b) What are the mole ratios of the product to each reactant and of the reactants to each other?

**Reactant to each other:**

4 mol Fe: 3 mol O<sub>2</sub>

**Product to each reactant:**

2 mol Fe<sub>2</sub>O<sub>3</sub>: 4 mol Fe

# Chapter 1 stoichiometry Practice Problems

**Edward I. Peters**



## **Chapter 1 stoichiometry Practice Problems:**

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**Fluid and Thermal Sciences** Nuggehally S. Nandagopal, PE, 2022-04-05 This text provides a clear understanding of the fundamental principles of thermal and fluid sciences in a concise manner in a rigorous yet easy to follow language and presentation Elucidation of the principles is further reinforced by examples and practice problems with detailed solutions Firmly grounded in the fundamentals the book maximizes readers capacity to take on new problems and challenges in the field of fluid and thermal sciences with confidence and conviction Standing also as a ready reference and review of the essential theories and their applications in fluid and thermal sciences the book is applicable for undergraduate mechanical and chemical engineering students students in engineering technology programs as well as practicing engineers preparing for the engineering license exams FE and PE in USA and abroad Explains the concepts and theory with a practical approach that readers can easily absorb Provides the just the right amount of theoretical and mathematical background needed making it less intimidating for the reader Covers fluid and thermal sciences in a straight forward yet comprehensive manner facilitating a good understanding of the subject matter Includes a wide spectrum and variety of problems along with numerous illustrative solved examples and many practice problems with solutions

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subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book CliffsAP Chemistry, 4th Edition Bobrow Test Preparation Services, 2011-09-26 Your complete guide to a higher score on the AP Chemistry exam Why CliffsAP Guides Go with the name you know and trust Get the information you need fast Written by test prep specialists Contents include Introduction overview of the test and how it is scored proven strategies for each type of question Review of topics tested atom periodic table bonding geometry hybridization stoichiometry gases liquids and solids thermodynamics solutions equilibrium acids and bases kinetics redox nuclear chemistry organic chemistry and writing reactions The Labs feature 20 multiple choice questions multiple free response questions on each topic with answers on each topic with answers and explanations scoring rubrics and 2 full length practice exams Structured like the actual exam Complete with answers and explanations AP is a registered trademark of the College Board which was not involved in the production of and does not endorse this product **Holt Chemistry** Ralph Thomas Myers, 2004 A Visual Analogy Guide to Chemistry, 2e Paul A Krieger, 2018-02-01 A Visual Analogy Guide to Chemistry is the latest in the innovative and widely used series of books by Paul Krieger This study guide delivers a big picture view of difficult concepts and effective study tools to help students learn and understand the details of general organic and biochemistry topics A Visual Analogy Guide to Chemistry is a worthwhile investment for any introductory chemistry student **McGraw-Hill's SAT Subject Test Chemistry, 3rd Edition** Thomas Evangelist, 2011-02-03 Expert guidance on the Chemistry exam Many colleges and universities require you to take one or more SAT II Subject Tests to demonstrate your mastery of specific high school subjects McGraw Hill's SAT Subject Test Chemistry is written by experts in the field and gives you the guidance you need perform at your best This book includes 4 full length sample tests updated for the latest test format 40 top tips to remember on test day Glossary of tested chemistry terms and formulas Tips and strategies from one of the most popular teachers at the renowned Brooklyn Technical High School Diagnostic test to pinpoint strengths and weaknesses Step by step review of all topics covered on the exam In depth coverage of the lab experiment questions that are a major test feature Charts tables and illustrations to simplify and reinforce learning Practice tests just like the real SAT Subject Test in Chemistry Test taking tips and strategies

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recommended tips and strategies to help you raise your score      **Schaum's Outline of College Chemistry, Ninth Edition** Jerome Rosenberg, Lawrence Epstein, Peter Krieger, 2009-10-16 Confusing Textbooks Missed Lectures Not Enough Time Fortunately for you there's Schaum's Outlines More than 40 million students have trusted Schaum's to help them succeed in the classroom and on exams Schaum's is the key to faster learning and higher grades in every subject Each Outline presents all the essential course information in an easy to follow topic by topic format You also get hundreds of examples solved problems and practice exercises to test your skills This Schaum's Outline gives you Practice problems with full explanations that reinforce knowledge Coverage of the most up to date developments in your course field In depth review of practices and applications Fully compatible with your classroom text Schaum's highlights all the important facts you need to know Use Schaum's to shorten your study time and get your best test scores Schaum's Outlines Problem Solved      **SAT Subject Test Chemistry** Joseph A. Mascetta, Mark Kernion, 2020-12-01 The updated edition of Barron's SAT Subject Test Chemistry includes A full length diagnostic test with explained answers Four practice tests that reflect the actual SAT Subject Test Chemistry All questions answered and explained Detailed reviews covering all test topics Appendixes which include the Periodic Table important equation constant and data tables and a glossary of chemistry terms Both teachers and test taking students have praised earlier editions of this manual for its wealth of well organized detail Subject reviewed include the basics matter energy scientific method and measurements atomic structure and the periodic table bonding chemical formulas gases and laws stoichiometry liquids solids and phase changes chemical reactions and thermochemistry chemical reactions chemical equilibrium acids bases and salts oxidation reduction carbon and organic chemistry and the laboratory ONLINE PRACTICE TESTS Students who purchase this book or package will also get access to two additional full length online SAT Chemistry subject tests with all questions answered and explained      **Problem Solving Guide and Workbook for Introductory Chemistry by Steve Russo, Mike Silver** Sandra Yancy McGuire, Steve Russo, Mike Silver, 2002 Provides over 175 worked examples and more than 500 practice problems and quiz questions to help students develop and practice their problem solving skills      **McGraw-Hill Education SAT Subject Test Chemistry, Fifth Edition** Thomas A. Evangelist, 2018-11-23 We Will Help You Get Your Best Score With more than 125 years of experience in education McGraw Hill Education is the name you trust to deliver results This MHE guide is the most comprehensive and relevant SAT Subject Test prep tool on the market This edition provides 5 full length practice tests with thorough answer explanations A comprehensive review of all Chemistry concepts essential to success on the SAT Subject Test An extensive overview of the format of the test based on the most recent SAT Chemistry exams Unique test taking strategies and tips recommended by teachers to help you raise your score A customizable study plan to help you maximize the time you have to prepare TOP 40 LIST The book includes a description of the 40 topics that are most crucial to know before you take the Subject Test in Chemistry TEST TAKING STRATEGIES Learn unique tips developed by teachers to help you avoid the test maker's traps

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web introduction this laboratory exercise relies on a titration technique to determine an unknown concentration of monoprotic acid in solution in the process of titration a basic solution is gradually added to the acidic solution until complete neutralization is obtained

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web chem 1210 spring 2019 experiment 10.11 part 1 acid base titration abstract the purpose of this experiment is to observe the titration of hydrochloric acid a strong acid with sodium hydroxide a strong base and acetic acid a

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1  $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$  one mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web jul 19 2023 introduction titration is an analytical quantitative technique used to determine the concentration of a solute a ph titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak pka of an unknown acid or pkb of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance  $\text{H}_2\text{C}_2\text{O}_4$  aq  $2\text{NaOH}$  aq  $\text{Na}_2\text{C}_2\text{O}_4$   $4\text{H}_2\text{O}$  to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2m hcl with the naoh of unknown concentration in order to experimentally ascertain the concentration of the naoh

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction  $\text{NaOH}$  aq  $\text{C}_6\text{H}_4\text{COOH}$  aq  $\text{C}_6\text{H}_4\text{COONa}$  aq  $\text{H}_2\text{O}$  l the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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