

Chapter 21 -- Nuclear Chemistry

21.1 Radioactivity

- When nuclei change spontaneously, emitting energy, they are said to be **radioactive**.
- **Nuclear chemistry** is the study of nuclear reactions and their uses.
- *Isotopes* have the same number of p⁺ but different numbers of n⁰.
 - Different isotopes of the same element are distinguished by their mass numbers.
 - Different isotopes have different natural abundances.
- A **radionuclide** is a radioactive nucleus.
 - Atoms containing these nuclei are called **radioisotopes**.

Nuclear Equations

- Most nuclei are stable.
 - Radionuclides are unstable and spontaneously emit particles and/or electromagnetic radiation.
 - Example: Uranium-238 is radioactive.
 - It emits **alpha (α) particles** (helium-4 nuclei)
- When a nucleus spontaneously decomposes in this manner, we say it has decayed (*radioactive decay*).
- In nuclear equations, the total number of nucleons is conserved.
 - We can represent the uranium-238 decay by the following nuclear equation:



- The total number of protons and neutrons before a nuclear reaction must be the same as the total number of nucleons after the reaction.

Types of Radioactive Decay

- There are three types of radiation which we will consider:
 - α-Radiation is the loss of ${}_2^4\text{He}$ (**alpha particles**) from the nucleus.
 - β-Radiation is the loss of an electron from the nucleus. These high-speed electrons are called **beta particles**.
 - γ-Radiation is the loss of high-energy photons from the nucleus.
- Nucleons can undergo two other types of decay:
 - **positron emission**
 - A positron is a particle with the same mass as an electron but with the opposite sign.
 - **electron capture**
 - The nucleus captures an electron from the electron cloud surrounding the nucleus.
- Representations:
 - In nuclear chemistry, to ensure the conservation of nucleons we write all particles with their atomic and mass numbers: ${}_2^4\text{He}$ and ${}_{-1}^0\text{e}$ represent α- and β- radiation.
- Nucleons can undergo decay:
 - (β-emission) ${}_0^1\text{n} \rightarrow {}_1^1\text{p} + {}_{-1}^0\text{e}$
 - (positron annihilation) ${}_1^0\text{e} + {}_{-1}^0\text{e} \rightarrow 2{}_0^0\gamma$
 - (positron or β⁺-emission) ${}_1^1\text{p} \rightarrow {}_0^1\text{n} + {}_1^0\text{e}$
 - (electron capture) ${}_1^1\text{p} + {}_{-1}^0\text{e} \rightarrow {}_0^1\text{n}$

Chapter 21 Nuclear Chemistry Section 3 Answers

E Durkheim



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Radiochemistry and Nuclear Chemistry Gregory Choppin, Jan-Olov Liljenzin, Jan Rydberg, Christian Ekberg, 2013-09-05
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Chemistry & Chemical Reactivity John C. Kotz,Paul Treichel,Patrick A. Harman,2003 Provides a broad overview of the principles of chemistry the reactivity of chemical elements and their compounds and the applications of chemistry Conveys a sense of chemistry as a field that not only has a lively history but also one that is currently dynamic with important new developments on the horizon **Organic Chemistry** Norman L. Allinger,1971 **The British National Bibliography Cumulated Subject Catalogue** ,1960

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