

Now _____ Date _____ Class _____

CHAPTER 9 **STUDY GUIDE FOR CONTENT MASTERY**

Covalent Bonding

Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	p bond
---------------	----------	------------	------------	--------

When sharing of electron occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5) p bond**.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is what form the elements such as hydrogen, nitrogen, and oxygen normally exist?
 - as single atoms
 - ☒ as molecules containing three atoms
 - as molecules containing two atoms
 - as molecules containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - four
- Bond length is the distance between
 - two molecules of the same substance
 - ☒ the nuclei of two attached atoms
 - the electrons in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - ☒ the shorter the bond, the stronger the bond
 - the shorter the bond, the weaker the bond
 - the shorter the bond, the lower the electron in it
 - the shorter the bond, the lower the bond dissociation energy

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Chemistry: Matter and Change • Chapter 9

49

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CHAPTER 9 **STUDY GUIDE FOR CONTENT MASTERY**

Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

- Binary molecular compounds are generally composed of a metal and a nonmetal. **false**
- The second element in the formula of a binary compound is named using the suffix -ite. **false**
- The prefix zero indicates three atoms. **false**
- The prefix four indicates six atoms. **true**
- In naming the first element in a formula, the prefix mono- is not used. **true**
- For binary acids, the hydrogen part of the compound is named using the prefix hydro-. **true**
- An oxyacid contains only two elements. **false**
- If the name of the anion of an oxyacid ends in -ite, the acid name contains the suffix -ous. **false**

In your textbook, read about naming molecular compounds and oxyacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A

- CO
- NO₂
- H₂CO₃
- NO₂
- N₂O₅
- HNO₃
- HNO₂
- HBr
- HBrO₃

Column B

- hydrobromic acid
- dinitrogen tetroxide
- carbon monoxide
- nitrous acid
- ammonia
- nitric acid
- carbonic acid
- bromic acid
- carbon dioxide

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50

Chapter 22 Study Guide For Content Mastery

Hydrocarbons Answer

Puja Mehta



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Table of Contents Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer

1. Understanding the eBook Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - The Rise of Digital Reading Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Personalized Recommendations
 - Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer User Reviews and Ratings
 - Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer and Bestseller Lists
5. Accessing Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer Free and Paid eBooks
 - Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer Public Domain eBooks
 - Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer eBook Subscription Services
 - Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer Budget-Friendly Options
6. Navigating Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer eBook Formats

- ePub, PDF, MOBI, and More
- Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer Compatibility with Devices
- Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Highlighting and Note-Taking Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Interactive Elements Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
- 8. Staying Engaged with Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
- 9. Balancing eBooks and Physical Books Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Setting Reading Goals Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Fact-Checking eBook Content of Chapter 22 Study Guide For Content Mastery Hydrocarbons Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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