Practice makes perfect with Foreday's Law Using conversion factors in Electrophesidatry

Determine the mass of copper that is plated out when a current of 10.0 acces is passed for 30.0 memotes through a solution containing Ua<sup>2\*</sup>

5-534 Answer

 How long will it take to plate out 1.0 kg Al from aqueous Al using a current of 100.0 A? Calculate in minutes

1000 A1 x imular x invit y 96 years x iscex lmin =

1288mw Answer 1,800 min

13. Calculate the moles of copper metal that can be produced by the dictrolysis of molten copper sulfate using \$00, C of electricity.

500 Cx \_ 1001= x \_ 1001 Cm

. 0 0 259 en 2:51 xce -3 meles

 What mass of copper could be deposited from a copper(III) suighable solution using a current of 0.50 A over 10 seconds?

.50 C x 10 5 x Inde x Inde x 6255}

1-6 X60 -3 9 ASSENT

5. Culculate the time required to deposit 56 g of silver from a silver nitrate rolution using a current of 4.5 A. — 41 — 0 (1)

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31 hrs

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Yan Bai

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