

CHEMICAL EQUILIBRIUM



EQUILIBRIUM CONSTANT 'K' •

For a General Reaction



The equilibrium constant expression is

$$K_c = \frac{[P]^p [Q]^q}{[A]^a [B]^b}$$

where K_c is the **Equilibrium Constant** (or K_p if they are all gases)

RELATION BETWEEN K_p AND K_c •

For the Reaction



$$K_p = \frac{[P_c]^c \times [P_D]^d}{[P_A]^a \times [P_B]^b} = \frac{[C]^c [D]^d}{[A]^a [B]^b} \frac{(RT)^{c+d}}{(RT)^{a+b}}$$

$$K_p = K_c (RT)^{\Delta n_g}$$

$$\text{if } \Delta n_g = 0 \implies K_p = K_c$$

Where, $\Delta n_g = (c+d) - (a+b)$

= no. of moles of gaseous products - no. of moles of gaseous Reactants

WHAT DOES THE VALUE OF 'K' MEAN ? •

$$\text{(a) } K \gg 1$$

Reactants



Products

if $K \gg 1$, the reaction is product-favoured; product predominates at equilibrium.

$$\text{(a) } K \ll 1$$

Reactants



Products

if $K \ll 1$, the reaction is reactant-favoured; reactant predominates at equilibrium.

$$\text{(a) } K \approx 1$$

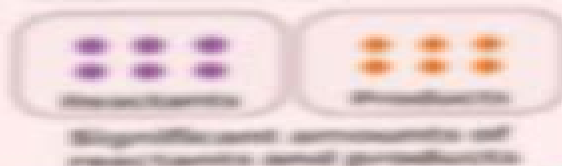
the reaction lies in the middle (no. of reactants and products)

MAGNITUDE OF 'K' •

Small ($K < 10^{-3}$)



Intermediate ($10^{-3} < K < 10^3$)



Large ($K > 10^3$)



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