

Conceptual Chemistry

PRACTICE PAGE

Chapter 10: Acids and Bases In Our Environment

Expert-Only Questions ♦♦♦

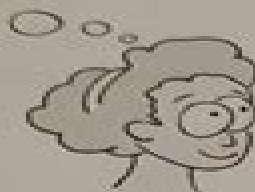
1. To say that an acid is strong in aqueous solution means that the acid
- produces a large concentration of hydronium ions.
 - is quite odorous when dissolved in water.
 - does not form chemical bonds with water.
 - does not separate into ions when dissolved in water.

$$10^3 = 1000$$

$$10^2 = 100$$

$$10^1 = 10$$

$$10^0 = ?$$



2. What does it mean to say that a solution has a pH of zero?
- There are no hydronium ions in that solution.
 - The concentration of hydronium ions is 1 M.
 - The solution is very alkaline.
 - It is not possible to have a pH equal to zero!

3. We measure the "acidity" of a solution by measuring the concentration of hydronium ions, H_3O^+ . What does it mean to say that one has a "neutral" solution, such as one has with pure water?

- There are no hydronium ions in this solution.
- The concentration of hydronium ions is very small.
- The concentration of hydronium ions equals the concentration of hydroxide ions, OH^- .



Pure?



4. Water reacting with itself is an endothermic reaction:



What happens, therefore, to the pH of neutral water as its temperature is increased and thus more thermal energy becomes available? The pH of neutral water

- remains at 7 because the pH of neutral water is always 7.
- decreases to some value below 7.
- increases to some value above 7.

The reaction of water with itself is favored by greater temperatures and pressures. At the super-high temperature and pressure of 374°C and 218 atmospheres, the reaction of water with water is so great that K_w equals 10^{-13} , which is ten times greater than the usual value of 10^{-14} . What is the pH of such a solution? Is this pH acidic, neutral, or basic? What might you predict about the corrosive properties of water under these conditions? So much to learn about our universe!



S. Subak

Conceptual Chemistry Practice Page Answers

Daniela Niemeyer



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