

## Molecular Models Lab Covalent Compounds

### Introduction

The purpose of this lab is to evaluate the geometries and polarities of simple covalent molecules. Many physical and chemical properties of compounds depend upon the shape or geometry of the molecule and its polarity. A hemoglobin molecule, for example, becomes ineffective as an oxygen transport vehicle if its shape is altered. Molecular shape determines a compound's boiling point, freezing point, and viscosity. A molecule's polarity determines whether it dissolves in water or not.

### Molecular Geometries

The geometry of a small molecule can be predicted by examining the central atom. The valence electron pairs that surround the central atom will repel each other. The electron pairs will position themselves so that there is a maximum distance between them. This rule is called the **VSEPR** theory (Valence Shell Electron Pair Repulsion). The electron pairs include the covalent bonds (shared pairs) and the unshared pairs of electrons. The 5 simplest geometries and their characteristics are found in the following table:

No. of Electron Pairs on Central Atom	Geometry Name	Central Atom Characteristics	Comments
No central atom	Linear	$X - X$	If a molecule only contains 2 atoms it is linear.
2 shared, 0 unshared	Linear	$— X —$ or $= X =$	Double and triple bonds only count as a shared pair for geometry purposes.
3 shared, 0 unshared	Trigonal planar	$\begin{array}{c}   \\ — X — \end{array}$ or $\begin{array}{c}   \\ = X = \end{array}$	Bond angles are $120^\circ$
4 shared, 0 unshared	Tetrahedral	$\begin{array}{c}   \\ — X — \\   \end{array}$	Since the molecule is 3-dimensional, the bond angles are $109^\circ$ , not $90^\circ$ .
2 shared, 2 unshared	Bent	$\begin{array}{c}   \\ : X : \\   \end{array}$	The bond angle is slightly $< 109^\circ$ , not $90^\circ$ .
3 shared, 1 unshared	Pyramidal	$\begin{array}{c}   \\ — X : \\   \end{array}$	The bond angle is slightly $< 109^\circ$ , not $90^\circ$ .

# Conclusion For Molecular Models Lab In Chemistry

**John W. Lehman**



## **Conclusion For Molecular Models Lab In Chemistry:**

**Chemistry (Teacher Guide)** Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter, helps each student stay on schedule and be organized, and is their source of accountability along the way. With that in mind, this guide provides additional help through the laboratory exercises as well as lessons, quizzes, and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study.

**Features**  
Each suggested weekly schedule has five easy-to-manage lessons that combine reading and worksheets. Worksheets, quizzes, and tests are perforated, and three-hole punched materials are easy to tear out, hand out, grade, and store. Adjust the schedule and materials needed to best work within your educational program. Space is given for assignments, dates. There is flexibility in scheduling. Adapt the days to your school schedule.

**Workflow**  
Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade.

**About the Author**  
DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and the Sable Institute of Environmental Studies.

*Exploring General Chemistry in the Laboratory* Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two-semester general chemistry course. The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life. This lab manual covers topics such as composition of compounds, reactivity, stoichiometry, limiting reactants, gas laws, calorimetry, periodic trends, molecular structure, spectroscopy, kinetics, equilibria, thermodynamics, electrochemistry, intermolecular forces, solutions, and coordination complexes. By the end of this course, you should have a solid understanding of the basic concepts of chemistry, which will give you confidence as you embark on your career in science.

Computational Modelling and Simulations for Designing of Corrosion Inhibitors Dakeshwar Kumar Verma, Chandrabhan Verma, Jeenat Aslam, 2023-04-19 Computational Modeling and

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