

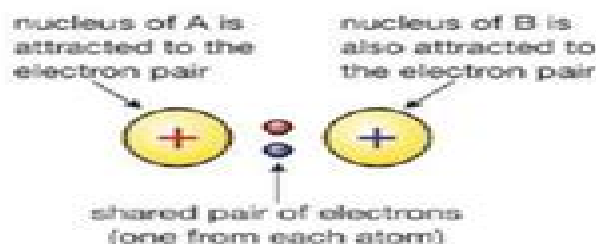
## CHAPTER 8: COVALENT BONDING

**1.44** Know that a covalent bond is formed between atoms by the sharing of a pair of electrons

- **Covalent Bond:** bond between non-metal and non-metal formed by a shared pair of electrons between the two atoms. It occurs between atoms of non-metal elements.

**1.45** Understand covalent bonds in terms of electrostatic attractions

- In a **covalent bond**, a pair of electrons is shared between two atoms. What holds the atoms together is the strong electrostatic attraction between the nuclei (positively charged) of the atoms that make up the bond, and the shared pair of electrons (negatively charged).



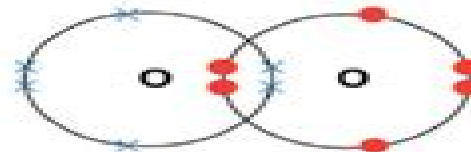
**1.46** Understand how to use dot-and-cross diagrams to represent covalent bonds in:

- Diatomic molecules, including hydrogen, oxygen, nitrogen, halogens and hydrogen halides

**Hydrogen ( $H_2$ )**



**Oxygen ( $O_2$ )**



# Core Teaching Resources Chapter 8 Covalent Bonding

**Julia Schneider**



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