

The Mole



A pair is of something.

A dozen is of something.

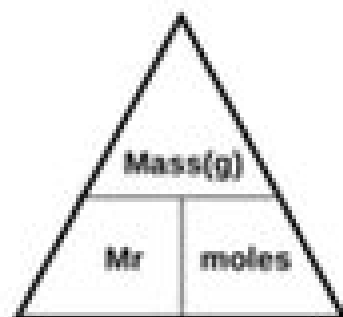
A mole is of something

A mole of water molecules has a mass of g, a mole of chlorine molecule has a mass of g.

The number 602 204 500 000 000 000 000 was carefully chosen so that the mass of that number of particles equals the *r*..... *a*..... mass (*Mr*) in g.....

For example, the *Mr* of water is 18 and so the mass of 1 mole of water molecules is g

If 1 mole of water has the mass of 18 g, then 2 moles has the mass ofg, the general equation is shown below. This can be remembered by thinking of 'Mr Moles'.



$$\text{Mass (g)} = \text{Mr} \times \text{moles}$$

Substance	Formula	Mr	Mass(g)	Moles
Water			36	$36/18=2$
Calcium carbonate			1 million g	
Ammonium nitrate	NH_4NO_3			150
Unknown gas	Unknown		3.2	0.10

1. How many moles are there in each of the following substances?

a) 8 g of SO_2

b) 371 g of Na_2CO_3

c) 1 Kg of Fe_2O_3

2. What mass would the following quantities have?

a) 2.5 moles of Cl_2

b) 20 moles of NH_4OH

c) 0.02 moles of $\text{Mg}(\text{NO}_3)_2$

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