

# Chapter 6

## Chapter 6: Thermochemistry

### 6.2 - Enthalpy and Calorimetry

Enthalpy

$$H = E + PV$$

For a chemical reaction, enthalpy change is given by the equation

$$\Delta H = H_{\text{products}} - H_{\text{reactants}}$$

Products of a reaction have a greater enthalpy ( $\Delta H$ ):

- $\Delta H$  is positive
- heat is absorbed by the system (endothermic reaction)

Reactants of a reaction have a greater enthalpy ( $\Delta H$ ):

- $\Delta H$  is negative
- heat is released by the system (exothermic reaction)

Calorimetry

Constant-pressure calorimetry

change in enthalpy = change in heat

$$\Delta H = q_p$$

# Core Teaching Resources Chemistry Chapter 6

**R Sanford**



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