

1. Show the electron configurations of all the electrons that show the electron configurations of the highest-energy electron using the quantum number technique for each of the following atoms.

(a) carbon $1s^2, 2s^2, 2p^2$

Quantum # $\begin{matrix} n & l \\ 2 & 1 \end{matrix}$

(b) Chromium⁺³ $1s^2, 2s^2, 3s^2, 3p^4, 4s^2, 3d^3$

Quantum $\begin{matrix} n & l \\ 3 & 2 \end{matrix}$

(c) Ca⁺² $1s^2, 2s^2, 2p^6, 3s^2$

Quantum # $\begin{matrix} n & l \\ 3 & 0 \end{matrix}$

2. Calculate the energy needed to cause an electron to jump from the 3rd shell to the 2nd shell. Energy =

$$E = \Delta E_{\frac{1}{n_1}} = \frac{1}{n_1^2} \frac{1}{2} = 2.18 \cdot 10^{-18} J \frac{1}{4} - \frac{1}{n_2^2} \frac{1}{2} = 2.18 \cdot 10^{-18} J \frac{1}{16} - \frac{1}{4} \frac{1}{16} = 2.18 \cdot 10^{-18} J$$

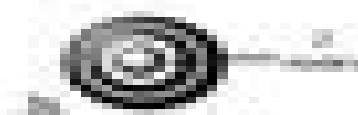
3. Calculate the wavelength for the photon that will be emitted when an electron falls spontaneously to the 2nd shell of a hydrogen atom. Wavelength

$$E = h\nu = \frac{hc}{\lambda} \quad \lambda = \frac{hc}{E} \quad \text{from these we get}$$

$$\lambda = \frac{hc}{\frac{1}{2} \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)} = 6.626 \cdot 10^{-34} \text{ J} \cdot \text{s} \cdot 2.99 \cdot 10^8 \text{ m/s} = 2.74 \cdot 10^{-7} \text{ m}$$

4. What does the solution to the Schrodinger equation for an electron represent?
 The solution to the Schrodinger equation represents the energy of an electron in an atomic orbital described by the wavefunction Ψ .
 $\hat{H}\Psi = E\Psi$

5. Draw the shape of the following orbitals.



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M Mark

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Table of Contents Core Teaching Resources Chemistry Chapter 7 Answer Key

1. Understanding the eBook Core Teaching Resources Chemistry Chapter 7 Answer Key
 - The Rise of Digital Reading Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Core Teaching Resources Chemistry Chapter 7 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Personalized Recommendations
 - Core Teaching Resources Chemistry Chapter 7 Answer Key User Reviews and Ratings
 - Core Teaching Resources Chemistry Chapter 7 Answer Key and Bestseller Lists
5. Accessing Core Teaching Resources Chemistry Chapter 7 Answer Key Free and Paid eBooks
 - Core Teaching Resources Chemistry Chapter 7 Answer Key Public Domain eBooks
 - Core Teaching Resources Chemistry Chapter 7 Answer Key eBook Subscription Services
 - Core Teaching Resources Chemistry Chapter 7 Answer Key Budget-Friendly Options
6. Navigating Core Teaching Resources Chemistry Chapter 7 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Core Teaching Resources Chemistry Chapter 7 Answer Key Compatibility with Devices
 - Core Teaching Resources Chemistry Chapter 7 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Highlighting and Note-Taking Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Interactive Elements Core Teaching Resources Chemistry Chapter 7 Answer Key

8. Staying Engaged with Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Core Teaching Resources Chemistry Chapter 7 Answer Key
9. Balancing eBooks and Physical Books Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Core Teaching Resources Chemistry Chapter 7 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Setting Reading Goals Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Fact-Checking eBook Content of Core Teaching Resources Chemistry Chapter 7 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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