

**10****CHEMICAL QUANTITIES****Chapter Test B****A. Matching**

Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line.

**Column A**

- \_\_\_\_\_ 1. the percent by mass of each element in a compound
- \_\_\_\_\_ 2. 0°C and 101.3 kPa or 1 atmosphere
- \_\_\_\_\_ 3. the lowest whole-number ratio of the atoms of elements in a compound
- \_\_\_\_\_ 4. the species present in a substance—usually atoms, molecules, or formula units
- \_\_\_\_\_ 5.  $6.02 \times 10^{23}$  representative particles of a substance
- \_\_\_\_\_ 6. the mass of an atom of an element
- \_\_\_\_\_ 7. the mass of one mole of any element or compound
- \_\_\_\_\_ 8. 22.4 L of any gas measured at STP

**Column B**

- a. molar mass
- b. empirical formula
- c. atomic mass
- d. molar volume
- e. representative particle
- f. percent composition
- g. one mole
- h. standard temperature and pressure

**B. Multiple Choice**

Choose the best answer and write its letter on the line.

- \_\_\_\_\_ 9. How many oxygen atoms are in 10 formula units of  $\text{Al}_2(\text{SO}_4)_3$ ?
  - a. 10 atoms O
  - b. 7 atoms O
  - c. 70 atoms O
  - d. 120 atoms O
- \_\_\_\_\_ 10. How many molecules are in 4.50 moles of  $\text{H}_2\text{O}$ ?
  - a. 450 molecules
  - b.  $2.71 \times 10^{28}$  molecules
  - c.  $6.02 \times 10^{23}$  molecules
  - d. 3.00 molecules
- \_\_\_\_\_ 11. How many moles are in  $8.5 \times 10^{25}$  molecules of  $\text{CO}_2$ ?
  - a.  $1.4 \times 10^2$  mol
  - b.  $7.1 \times 10^{-25}$  mol
  - c.  $5.1 \times 10^{19}$  mol
  - d.  $8.5 \times 10^{25}$  mol

# Core Teaching Resources Chemistry Test Answer Key

**D Siedentop**



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