

ANSWER KEY to Chemistry: Dimensional Analysis – More Practice Problems

1. Determine the number of mm in 1600 m.

$$\frac{1600 \text{ m}}{1} \times \frac{1000 \text{ mm}}{1 \text{ m}} = 1,600,000 \text{ mm or } 1.6 \times 10^6 \text{ mm}$$

2. Determine the number of m in 1600 mm.

$$\frac{1600 \text{ mm}}{1} \times \frac{1 \text{ m}}{1000 \text{ mm}} = 1.6 \text{ m}$$

3. Determine the number of mm in 14.3 cm.

$$\frac{14.3 \text{ cm}}{1} \times \frac{10 \text{ mm}}{1 \text{ cm}} = 143 \text{ mm}$$

4. Convert 5.2 cm of magnesium (Mg) ribbon to mm of Mg ribbon.

$$\frac{5.2 \text{ cm}}{1} \times \frac{10 \text{ mm}}{1 \text{ cm}} = 52 \text{ mm}$$

5. Convert 0.049 kg sulfur (S) to g of S.

$$\frac{0.049 \text{ kg}}{1} \times \frac{1000 \text{ g}}{1 \text{ kg}} = 49 \text{ g}$$

6. Convert 150 cm of copper (Cu) wire into mm of Cu wire.

$$\frac{150 \text{ cm}}{1} \times \frac{10 \text{ mm}}{1 \text{ cm}} = 1500 \text{ mm}$$

7. Convert 150 mg of acetylsalicylic acid (aspirin) to g of aspirin.

$$\frac{150 \text{ mg}}{1} \times \frac{1 \text{ g}}{1000 \text{ mg}} = 0.150 \text{ g}$$

8. Convert 2500 mL of hydrochloric acid (HCl) to L of HCl.

$$\frac{2500 \text{ mL}}{1} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 2.5 \text{ L}$$

9. Convert 0.020 kg of tin (Sn) to mg of Sn.

$$\frac{0.020 \text{ kg}}{1} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1000 \text{ mg}}{1 \text{ g}} = 20,000 \text{ mg}$$

10. Convert 2875 cm³ to liters.

$$\frac{2875 \text{ cm}^3}{1} \times \frac{1 \text{ mL}}{1 \text{ cm}^3} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 2.875 \text{ L}$$

11. The density of lead (Pb) is 11.34 g/mL. Find the density of Pb in kg/L.

$$\frac{11.34 \text{ g}}{1 \text{ mL}} \times \frac{1 \text{ kg}}{1000 \text{ g}} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.00001134 \text{ kg/L or } 1.134 \times 10^{-5} \text{ kg/L}$$

12. Convert 0.5 g of sodium (Na) to kg of Na.

$$\frac{0.5 \text{ g}}{1} \times \frac{1 \text{ kg}}{1000 \text{ g}} = 0.0005 \text{ kg or } 5 \times 10^{-4} \text{ kg}$$

13. How many seconds are in 4.3 years?

$$\frac{4.3 \text{ yrs}}{1} \times \frac{365.25 \text{ days}}{1 \text{ yr}} \times \frac{24 \text{ hrs}}{1 \text{ day}} \times \frac{60 \text{ min}}{1 \text{ hr}} \times \frac{60 \text{ sec}}{1 \text{ min}} = 135697680 \text{ sec or } 1.40000000 \times 10^8 \text{ sec or } 1.4 \times 10^8 \text{ sec}$$

14. A metallurgist is making an alloy that consists of 325 g of chromium (Cr) and 2.5 kg of iron (Fe). Find the total mass of the mixture in kg.

$$\frac{325 \text{ g}}{1} \times \frac{1 \text{ kg}}{1000 \text{ g}} = 0.325 \text{ kg} \quad 0.325 \text{ kg} + 2.5 \text{ kg} = 2.8 \text{ kg (notice sig figs!)}$$

15. How many mL of water (H₂O) will it take to fill a 2.0 L bottle that already contains 1.87 L of H₂O?

$$2.0 \text{ L} - 1.87 \text{ L} = 0.13 \text{ L}$$

$$\frac{0.13 \text{ L}}{1} \times \frac{1000 \text{ mL}}{1 \text{ L}} = 130 \text{ mL}$$

Note: technically, you should round the 0.13 L to 0.1 L... because of sig figs... which then makes the next calculation = 100 mL.

Dimensional Analysis Practice Chemistry Answers

Mary Millhollon, Richard H. Langley



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