

Section 1.3: The Carbon Chemistry of Life

Section 1.3 Questions, page 28

1. Carbon has four unpaired valence electrons allowing it to form four covalent bonds. Carbon can bond with one or two other carbons to form straight chains. When carbon bonds with three or four other carbons, each secondary carbon can form a chain resulting in a branched chain. Due to electron repulsion, the bonds form a tetrahedral structure. The angles between the bonds allow chains of carbon to bend around and form ring structures.

2. Functional groups are specific groups of atoms within organic molecules that are responsible for the characteristic chemical reactions of those molecules. The same functional group will undergo the same or similar chemical reaction(s) regardless of the size of the molecule it is a part of.

3. Compounds with similar structures often have similar uses. Their use is dependent on the properties of the compounds and the properties are dependent on structures of the compounds.

4. Answers may vary. Sample answer: Functional groups help determine if a molecule is polar or non-polar. This property determines the types of solvents and molecules that it is attracted to. For example, polar solvents dissolve polar and ionic compounds and non-polar solvents dissolve non-polar compounds.

5. Answers may vary. Sample answer: In many functional groups, oxygen helps establish the polarity of the molecule. In some groups, oxygen acts as a reducing agent by losing electrons from compounds while in other groups, it acts as an oxidizing agent by gaining electrons to form a bond.

6. (a) $\text{-CO}_2\text{H} + \text{H}_2\text{O} \rightarrow \text{-CO}_2^- + \text{H}_3\text{O}^+$

The carboxyl group loses a hydrogen ion.

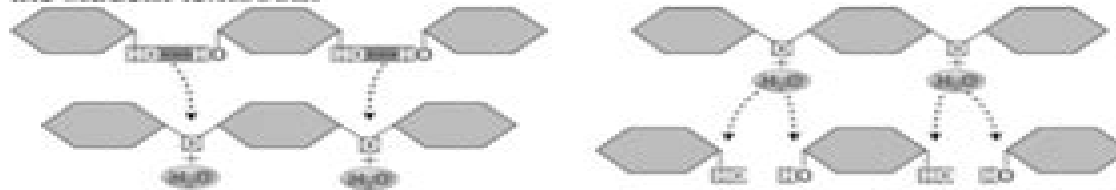
(b) $\text{-NH}_2 + \text{H}_2\text{O} \rightarrow \text{-NH}_3^+ + \text{OH}^-$

The amino group gains a hydrogen ion.

(c) $\text{-PO}_3\text{H}_2 + \text{OH}^- \rightarrow \text{-PO}_3\text{H}^- + \text{H}_2\text{O} \rightarrow \text{-PO}_3^{2-} + \text{H}_3\text{O}^+$

The phosphate group loses a hydrogen ion.

7. Answer may vary. Sample answer: Student drawings should resemble Figure 6 on page 27 of the student textbook.



(a)

(b)

8. Dehydration synthesis and hydrolysis are both processes that rearrange molecules in organic substances. Dehydration synthesis connects smaller molecules to form larger organic compounds. Water is lost during the process. In hydrolysis, water is added and larger compounds are broken down into smaller units.

9. This is an example of hydrolysis because the products are smaller than the reactants and water is added.

10. Answers may vary. Sample answer: Dehydration could be used to form longer and more complex polysaccharides for cell function. Hydrolysis could be used to reduce polysaccharides into simple more easily utilized sugars.

Carbon Chemistry Wordwise Answer Key

RJ Alexander



Carbon Chemistry Wordwise Answer Key:

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Carbon Chemistry, 1991 **Key to Carbon Chemistry**, 1999

Carbon Chemistry Alpha & Omega Publishing, 2001-03-01 **Carbon Chemistry Student Text** Ellen McHenry, 2021-08-19 This is the student text section of the curriculum book Carbon Chemistry by the same author The text has nine chapters covering these topics structure of carbon atoms hydrocarbons alkanes alkenes alkynes functional groups polymers plastics soap rubber nylon carbon dioxide and monoxide limestone the carbon cycle radioactive carbon 15 This student text is for parents or teachers who will be working with more than one student and would like each student to have their own copy of the book **Carbon Chemistry** Krista West, 2009 Learn how carbon is integrated into all facets of life as we know it and discover the unique properties of this essential element P 4 of cover *ChemLab: Carbon chemistry* Brian J. Knapp, 1998 *Carbon Chemistry* Brian J. Knapp, 1998 Theories and concepts of carbon compounds with and without hydrogen are explained through the use of a series of illustrated step by step demonstrations *ChemLab*, 1997 Carbon Chemistry Lesley Love, Rod McLeod, Pamela Harvey, 1987 *Carbon Chemistry Teacher's Guide* Benchmark Education Co. LLC Staff, 2014-01-01 Single title not sold as a separate item Sold as part of larger package only **HSC Carbon Chemistry** C. M. Roebuck, 1989 Dynamics of Being Brown, S B, Ed, 1980 Glencoe Sci Chemistry Chapter 4 Carbon Chemistry Chp Res 652 2002 Alton Biggs, McGraw-Hill Staff, National Geographic Society

(U.S.), Glencoe/McGraw-Hill, 2001-08-01 **An introduction to the carbon compounds** Richard Haliburton Adie, 1899

Chemistry of the Carbon Compounds; Or, Organic Chemistry, 1886 Chemistry of Carbon Compounds, 1962

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ω represents the frequency of a sine wave when we write it this way $\sin \omega t$ if $\omega = 1$ the sin completes one cycle in 2π seconds if $\omega = 2\pi$ the sin completes one cycle sooner every 1 second for $f = 1$ if $\omega = 2\pi f$ the sin completes one cycle even faster every $1/f$ seconds or equivalently sin completes f cycles in 1 second

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