Name \_\_\_\_

#### Solution Stoichiometry Worksheet

Solve the following solutions Stoichiometry problems:

 How many grams of silver chromate will precipitate when 150, mL of 0,500 M silver nitrate are added to 100, mL of 0,400 M potassium chromate?

2 AgNONet + KcCrOwed - AgcCrOwel + 2 KNOwed

0.150 L AgNO <sub>2</sub>	0.500 moles AgNO <sub>3</sub>	1 moles Ag <sub>2</sub> CrO <sub>4</sub>	331.74 g Ag:CrO <sub>4</sub> = 13	2.4 g Age CrO.
	1 L	2 moles AgNO <sub>3</sub>	1 moles Ag <sub>2</sub> CrO <sub>4</sub>	

0.100 L K <sub>2</sub> CrO <sub>4</sub>	0.400 moles K <sub>2</sub> CrO <sub>4</sub>	1 moles Ag <sub>2</sub> CrO <sub>4</sub>	331.74 g Ag-CrO.	= 13.3 g Ag: CrO.
	1 L	1 moles K <sub>2</sub> CrO <sub>4</sub>	1 moles Ag <sub>2</sub> CrO <sub>4</sub>	

 How many mL of 0.280 M barium nitrate are required to precipitate as barium sulfate all the sulfate ions from 25.0 mL of 0.350 M aluminum sulfate? (93.8 mL barium nitrate)

3 Ba(NOs)sect + Ale(SOs)sect - 3 BaSOsect + 2 Al(NOs)sect

0.0250 L Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	0.350 moles Al <sub>2</sub> (SO <sub>4</sub> ) <sub>5</sub>	3 moles Ba(NO <sub>3</sub> ) <sub>2</sub>	1 L	= 0.0938 L. Ba(NO <sub>3</sub> ) <sub>2</sub>
	I L	1 moles Al <sub>2</sub> (SO <sub>a</sub> ) <sub>3</sub>	0.280 moles Ba(NO <sub>3</sub> ) <sub>2</sub>	

 25.0 mL of 0.350 M NaOH are added to 45.0 mL of 0.125 M copper (II) sulfate. How many grams of copper (II) hydroxide will precipitate?

2 NaOH<sub>(ac)</sub> + CuSO<sub>total</sub> → Cu(OH)<sub>(bc)</sub> + Na<sub>5</sub>SO<sub>total</sub>

0.0250 L NaOH	0.350 moles NaOH	1 moles Cu(OH) <sub>2</sub>	97.57 g Cu(OH) <sub>2</sub>	= 0.427 g Cu(OH) <sub>2</sub>
	1 L NaOH	2 moles NaOH	1 mole Cu(OH) <sub>2</sub>	

0.0450 L CuSO <sub>4</sub>	0.125 moles CuSO,	1 moles Cu(OH) <sub>2</sub>	97.57 g Cu(OH) <sub>2</sub>	= 0.549 g Cu(OH) <sub>2</sub>
	1 L NaOH	1 moles CuSO.	1 mole Cu(OH) <sub>2</sub>	

4. What volume of 0.415 M silver nitrate will be required to precipitate as silver bromide all the bromide ion in 35.0 mL of 0.128 M calcium bromide?

2 AgNO<sub>3(ac)</sub> + CaBr<sub>3(ac)</sub> → Ca(NO<sub>3</sub>)<sub>3(ac)</sub> + 2 AgBr<sub>(c)</sub>

0.0350 L CaBr <sub>2</sub>	0.128 moles CaBr <sub>2</sub>	2 moles AgNO <sub>3</sub>	1 L AgNO <sub>3</sub>	= 0.0216 L AgNO <sub>3</sub>
	1 L CaBr <sub>2</sub>	1 moles CaBr <sub>2</sub>	0.415 mole AgNO <sub>a</sub>	7.5

What volume of 0.496 M HCl is required to neutralize 20.0 mL of 0.809 M sodium hydroxide?

HCl(aq) + NaOH(aq) → NaCl(aq) + H(OH)(I)

0.0200 L NaOH	0.809 mole NaOH	1 mole HCI	1 L HCI	= 0.0326 L HCI
	1 L NaOH	1 mole NaOH	0.496 mole HCI	

**D** Siedentop

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# Table of Contents Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe

- 1. Understanding the eBook Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - The Rise of Digital Reading Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - Personalized Recommendations
  - o Chapter 1 content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe User Reviews and Ratings
  - Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe and Bestseller Lists
- 5. Accessing Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe Free and Paid eBooks
  - Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe Public Domain eBooks

- Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe eBook Subscription Services
- Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe Budget-Friendly Options
- 6. Navigating Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe eBook Formats
  - o ePub, PDF, MOBI, and More
  - Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe Compatibility with Devices
  - Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - Highlighting and Note-Taking Chapter 1 content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - Interactive Elements Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
- 8. Staying Engaged with Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
- 9. Balancing eBooks and Physical Books Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - Setting Reading Goals Chapter 1 content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe

- Fact-Checking eBook Content of Chapter 1content Mastery Stoichiometry Worksheet Key Teacher S Edition Glencoe
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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