

**Guided Practice - Acids & Bases****Properties of Acids and Bases**

1. Match the following terms with their descriptions:

- |                        |       |  |
|------------------------|-------|--|
| a. Arrhenius Acid      | _____ | the substance formed when a proton is added to _____       |
| b. Arrhenius Base      | _____ | the remaining substance when a proton is lost from an acid |
| c. Conjugate Acid      | _____ | a substance that produces $\text{OH}^-$ in solution        |
| d. Conjugate Base      | _____ | a substance producing $\text{H}^+$ in solution             |
| e. Bronsted-Lowry Acid | _____ | a proton acceptor  |
| f. Bronsted-Lowry Base | _____ | a proton donor   |

2. Which of the following represent conjugate acid-base pairs? Label each species as a conjugate acid or a conjugate base. For those pairs that are  $\text{H}_2\text{O}$  conjugates, write the correct conjugate acid or base for each species in the pair, labeling each species correctly.

- $\text{HSO}_4^-$ ,  $\text{SO}_4^{2-}$
- $\text{HBr}$ ,  $\text{Br}^-$
- $\text{H}_3\text{PO}_4$ ,  $\text{PO}_4^{3-}$
- $\text{H}_2\text{O}$ ,  $\text{HO}^-$

3. Complete the following table regarding acid strength.

| Property  | Strong Acid | Weak Acid |
|---|-------------|-----------|
| Definition  |             |           |
| Describe Dissociation                                 |             |           |
| Species present in solution                           |             |           |
| Relative strength of conjugate base compared to water |             |           |
| Equation showing acid's reaction to water             |             |           |

# Chapter 1review Acids Bases Answer Key

**RJ Shavelson**



## **Chapter 1 review Acids Bases Answer Key:**

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