

Unit 08 – Stoichiometry – Worksheet 1

Answer the following questions in the space provided at the left.

- 18 mol O₂ 1. How many moles of oxygen are produced from twelve moles of Potassium chlorate?



$$\frac{12 \text{ mol } \cancel{\text{KClO}_3}}{\cancel{2 \text{ mol } \text{KClO}_3}} \quad | \quad \frac{3 \text{ mol O}_2}{2 \text{ mol } \cancel{\text{KClO}_3}}$$

- 21 mol O₂ 2. How many moles of oxygen are produced from 14 moles of Potassium chlorate?



$$\frac{14 \text{ mol } \cancel{\text{KClO}_3}}{\cancel{2 \text{ mol } \text{KClO}_3}} \quad | \quad \frac{3 \text{ mol O}_2}{2 \text{ mol } \cancel{\text{KClO}_3}}$$

- 45 mol H₂O 3. How many moles of water are produced from 22.5 moles of oxygen?



$$\frac{22.5 \text{ mol } \cancel{\text{O}_2}}{\cancel{1 \text{ mol } \text{O}_2}} \quad | \quad \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol } \cancel{\text{O}_2}}$$

- 20 mol HCl 4. Calculate the number of moles of Hydrogen chloride produced from 10 moles of hydrogen.



$$\frac{10 \text{ mol } \cancel{\text{H}_2}}{\cancel{1 \text{ mol } \text{H}_2}} \quad | \quad \frac{2 \text{ mol HCl}}{1 \text{ mol } \cancel{\text{H}_2}}$$

- 21 mol Cl₂ 5. Calculate the number of moles of chlorine needed to form 14 moles of Iron (III) chloride.



$$\frac{14 \text{ mol } \cancel{\text{FeCl}_3}}{\cancel{2 \text{ mol } \text{FeCl}_3}} \quad | \quad \frac{3 \text{ mol Cl}_2}{2 \text{ mol } \cancel{\text{FeCl}_3}}$$

Chapter 1supplemental Problems Stoichiometry Answer Key

McGraw-Hill Staff



Chapter 1 supplemental Problems Stoichiometry Answer Key:

Chemistry McGraw-Hill Staff, 2001-03 *Holt Chemistry* Ralph Thomas Myers, 2004
Set/student's Version for Jones's Organic Chemistry Timo Ovaska, 2000

Supplemental Problems

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