MOLE CONCEPT

CLASS Xth

- Calculate the relative molecular masses (or molecular weights) of the following compounds:
- Copper sulphate crystals, CuSO4.5H2O.
- Ammonium sulphate, (NH4)2SO4.
- iii) Cane sugar, C12H22O11.

(Given atomic masses Cu = 63.5, S = 32, O = 16, N = 14, C = 12)

- An organic compound with vapour density = 94 contains. C = 12.67%, H = 2.13%, and Br = 85.11%. Find the molecular formula. [Atomic mass : C = 12, H = 1, Br = 80]
- (i) The compound A has the following percentage composition by mass. Carbon = 26.7%, oxygen = 71.1%, hydrogen = 2.2%. Determine the Empirical formula of compound A [work to one decimal place] [H = 1, C = 12, O = 16].
- ii) If the relative molecular mass of A is 90, what is the molecular formula of A?

(iii) The compound A is weak acid. What is meant by this statement?

- Determine the empirical formula of the compound whose composition by mass is: 42% nitrogen, 48% oxygen and 9% hydrogen. [H = 1; N = 14; O = 16]
- Determine the empirical formula of a compound containing 47.9% potassium, 5.5% beryllium and 46.6% fluorine by mass. (Atomic weight of Be = 9; F = 19; K = 39) Work to one decimal place.
- The Empirical formula of a compound is C2H5. It has a vapour density of 29. Determine
 the relative molecular formula mass of the compound and hence its molecular formula.
- The percentage composition of sodium phosphate as determined by analysis, is 42.1% sodium, 18.9% phosphorus and 39% oxygen. Find the empirical formula of the compound (work to two decimal places).
- A gaseous organic compound contains 3.6 g of carbon and 0.8 g of hydrogen. The vapour density of this compound is 22. (i) Calculate the Empirical formula. (ii) Calculate the molecular formula of the compound.

Chapter 1the Mole Answers

David A. Ucko

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Table of Contents Chapter 1the Mole Answers

- 1. Understanding the eBook Chapter 1the Mole Answers
 - The Rise of Digital Reading Chapter 1the Mole Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chapter 1the Mole Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 1the Mole Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chapter 1the Mole Answers

- Personalized Recommendations
- Chapter 1the Mole Answers User Reviews and Ratings
- Chapter 1the Mole Answers and Bestseller Lists
- 5. Accessing Chapter 1the Mole Answers Free and Paid eBooks
 - Chapter 1the Mole Answers Public Domain eBooks
 - Chapter 1the Mole Answers eBook Subscription Services
 - Chapter 1the Mole Answers Budget-Friendly Options
- 6. Navigating Chapter 1the Mole Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chapter 1the Mole Answers Compatibility with Devices
 - Chapter 1the Mole Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Chapter 1the Mole Answers
 - Highlighting and Note-Taking Chapter 1the Mole Answers
 - Interactive Elements Chapter 1the Mole Answers
- 8. Staying Engaged with Chapter 1the Mole Answers
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 1the Mole Answers
- 9. Balancing eBooks and Physical Books Chapter 1the Mole Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 1the Mole Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 1the Mole Answers
 - Setting Reading Goals Chapter 1the Mole Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 1the Mole Answers

- Fact-Checking eBook Content of Chapter 1the Mole Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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