

Chapter 2

I. Atoms and Molecules

- A. Emergent properties – atoms and molecules
- B. Life requires about 25 chemical elements
- C. Elements combine to form compounds - 2 or more elements in fixed ratio
 - 1. table salt is NaCl; iodized (fig. 2.4)
 - 2. Water
- D. Atoms consist of protons, neutrons and electrons
 - 1. atom is the smallest unit of matter
 - 2. subatomic particles make up the atom.
 - 3. Differences in elements
 - 4. isotopes - different numbers of neutrons so mass changes
- E. Electron arrangement determines chemical properties of atom
 - 1. electrons determine how an atom behaves
 - 2. electrons vary in energy
 - 3. electrons occur at certain energy levels
- F. Ionic bonds are attractions between ions of opposite charges
 - 1. ion = loss or gain of electron
 - 2. ions with opposite charges attract each other
 - 3. resulting compound is electrically neutral
 - 4. sodium and chloride ions are always present in 1:1 ratio
- G. Covalent bonds = sharing of electrons in outer shell
 - 1. molecule
 - 2. single covalent bond
 - 4. double bond
 - 5. H₂ and O₂ are molecules but not compounds
- H. Chemical reactions rearrange matter
 - 1. chemical reactions
 - 2. reactants
 - 3. product
 - 4. $2\text{H}_2 + \text{O}_2 = \text{H}_2\text{O}$

II. Properties of water

- A. water is a polar molecule
- B. Water's polarity leads to H bonding
- C. Hydrogen bonds make liquid water cohesive
 - 1. H bonds last for only a few trillionths of a second
 - 2. cohesion and adhesion
 - 3. surface tension
- D. Hydrogen bonds of water moderate temperature
 - 1. heat
 - 2. temperature
- E. Ice is less dense than water
- F. water is a versatile solvent
 - 1. solution
 - 2. solvent
 - 3. solute
 - 4. aqueous solution

III. pH Scale

- A. The chemistry of life is sensitive to acid and basic conditions
 - 1. some water molecules break apart naturally in water
 - 2. ions formed are called hydrogen (H⁺) and hydroxide (OH⁻) ions
 - 3. for life chemistry, proper balance of these ions is necessary
 - 4. the more acidic a solution, the more H⁺ ions. Acids are substances that donate H⁺
 - 5. in your stomach, HCl donates H⁺ ions when in solution

Chapter 2 Basic Chemistry Answers

Jianjun Gao



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