

Acids, Bases, and Salts Review

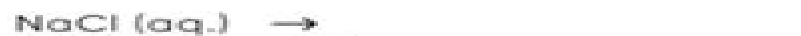
- ① An Arrhenius acid is defined as any compound dissociating in an aqueous solution to form _____ ions. Example: $\text{HNO}_3(\text{aq.}) \rightarrow \text{H}^+(\text{aq.}) + \text{NO}_3^-(\text{aq.})$. Complete the following equation:



- ② An Arrhenius base is defined as any compound dissociating in an aqueous solution to form _____ ions. Example: $\text{KOH}(\text{aq.}) \rightarrow \text{K}^+(\text{aq.}) + \text{OH}^-(\text{aq.})$. Complete the following equation:



- ③ Salts are compounds that dissociate in an aqueous solution releasing neither _____ nor _____ ions. Example: $\text{KCl}(\text{aq.}) \rightarrow \text{K}^+(\text{aq.}) + \text{Cl}^-(\text{aq.})$. Complete the following equation:



- ④ A solution that has strong conductivity and reacts with the metal to make hydrogen gas is a

☐ a Strong Base ☐ b Weak Base ☐ c Strong Acid ☐ d Weak Acid

Answers ____

- ⑤ Which of the following is slippery to touch?

☐ a Strong Acid ☐ b Strong Base ☐ c Strong Electrolyte ☐ d Water

Answers ____

- ⑥ The hydronium ion is

☐ a H_3O^+ ☐ b H_2O ☐ c OH^- ☐ d H_2

Answers ____

- ⑦ A substance that changes color in an acidic or basic solution is called

☐ a Neutralizer ☐ b Inhibitor ☐ c pH indicator ☐ d Catalyst

Answers ____

- ⑧ Name the following acids and bases

HF _____

HNO_3 _____

KOH _____

$\text{Mg}(\text{OH})_2$ _____

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Arnold B. Loebel



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web electrochemistry recommended mcqs 123 questions electrochemical series 8 qs 1 q no standard reduction electrode potential of three metals x y and z are 1 2 v 0 5 v and 3 v respectively the reducing power of these metals will be 1 y x z

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web and hold up short answers more complex questions could be discussed in groups not all answers to questions are included in the video some of the questions will draw on prior learning or extend learners thinking beyond the video content student version the same questions are offered as a printable worksheet for learners

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web this is the last lab of the semester and there is no formal lab report instead this is an exploratory lab where you will design an experiment using electrochemistry to determine the concentration of an unknown solution of copper ii

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web the electrochemical series is a list of elements ordered by their standard electrode potentials it tells us how easily each element is oxidised compared to a reference half cell the hydrogen electrode the electrochemical series is the basis behind all

kinds of modern fuel cells and batteries

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web introduction chemical reactions involving the transfer of electrons from one reactant to another are called oxidation reduction reactions or redox reactions in a redox reaction two half reactions occur one reactant gives up electrons undergoes oxidation and another reactant gains electrons undergoes reduction

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web oct 22 2019 electrochemical series is defined as the arrangement of the elements or their ions in increasing or decreasing order of their standard electrode potential under standard conditions electrode potential is also reduction potential and it is calculated for many elements by comparing with standard electrode potential of hydrogen electrode

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web an experiment that will illuminate the mysteries of ions create a cell of electrolyte and metals and watch how the reactive and non reactive metals form ion this practical helps students to grasp the ideas around electromotive force found in chemicals this experiment should take 45 minutes

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web the electrochemical series equilibrium e volts 3 03 2 92 2 87 2 71 2 37 1 66 0 76 0 44 0 13 0 0 34 0 80 1 50 a note on the hydrogen value remember that each e value shows whether the position of the equilibrium lies

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web this lab provides the opportunity to understand the concepts and set up of electrochemical reactions understanding electrochemistry is important for its practical applications electrochemical cells are used to purify metals produced by the processes of

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web sep 6 2023 complete step by step answer the electrochemical series is defined as the arrangement of the elements in their increasing order of electrode potential values the electrochemical series is sometimes referred to as activity series

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web electrochemical cell 1 select the following red wire metal silver solution 1 0 m agno 3 aq black wire metal copper solution 1 0 m cu no 3 2 aq label the diagram below showing the following electrochemical cell that you built and observed label electrode with the metals show direction of electron flow show electrode polarity

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web 100 7 ratings as per published electrochemical series the electrode potentials are cu 0 34v al 1 66v zn view the full answer transcribed image text list the electrochemical series you developed from this lab based on

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web choose 1 answer choice a at equilibrium e c e l l text cell cell start text end text start subscript c e l l end subscript calculates to a value of 0 and the battery is considered dead because q becomes k e q text eq eq

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web quickly measure what you know about the electrochemical series with this quiz and corresponding worksheet you must be familiar with such terms as electrolysis oxidation and reduction in order

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web this is a post lab for electrochemistry determining an activity series using galvanic cells these are the first 6 questions and this is my data but i only need answers for 7 and 8 1 using copper as the standard cu cu cell potential 0 determine the potential for each of the reactions between two metals