

Chapter 20 Worksheet: Redox Answers

R.A. = reducing agent
O.A. = oxidizing agent

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

- $2\text{Sr} + \text{O}_2 \longrightarrow 2\text{SrO}$ $\text{Sr}^0 \text{ to } \text{Sr}^{2+}$; oxidized / reducing agent $\text{O}^0 \text{ to } \text{O}^{2-}$; reduced / oxidizing agent
- $2\text{Li} + \text{S} \longrightarrow \text{Li}_2\text{S}$ $\text{Li}^0 \text{ to } \text{Li}^{+}$; oxidized / R.A. $\text{S}^0 \text{ to } \text{S}^{2-}$; reduced / O.A.
- $2\text{Cs} + \text{Br}_2 \longrightarrow 2\text{CsBr}$ $\text{Cs}^0 \text{ to } \text{Cs}^{+}$; oxidized / R.A. $\text{Br}^0 \text{ to } \text{Br}^{-}$; reduced / O.A.
- $3\text{Mg} + \text{N}_2 \longrightarrow \text{Mg}_3\text{N}_2$ $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$; oxidized / R.A. $\text{N}^0 \text{ to } \text{N}^{3-}$; reduced / O.A.
- $4\text{Fe} + 3\text{O}_2 \longrightarrow 2\text{Fe}_2\text{O}_3$ $\text{Fe}^0 \text{ to } \text{Fe}^{3+}$; oxidized / R.A. $\text{O}^0 \text{ to } \text{O}^{2-}$; reduced / O.A.
- $\text{Cl}_2 + 2\text{NaBr} \longrightarrow 2\text{NaCl} + \text{Br}_2$ $\text{Cl}^0 \text{ to } \text{Cl}^{-}$; reduced / O.A. $\text{Br}^{-} \text{ to } \text{Br}^0$; oxidized / R.A.
- $\text{Si} + 2\text{F}_2 \longrightarrow \text{SiF}_4$ $\text{Si}^0 \text{ to } \text{Si}^{4+}$; oxidized / R.A. $\text{F}^0 \text{ to } \text{F}^{-}$; reduced / O.A.
- $2\text{Ca} + \text{O}_2 \longrightarrow 2\text{CaO}$ $\text{Ca}^0 \text{ to } \text{Ca}^{2+}$; oxidized / R.A. $\text{O}^0 \text{ to } \text{O}^{2-}$; reduced / O.A.
- $\text{Mg} + 2\text{HCl} \longrightarrow \text{MgCl}_2 + \text{H}_2$ $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$; oxidized / R.A. $\text{H}^{+} \text{ to } \text{H}^0$; reduced / O.A.
- $2\text{Na} + 2\text{H}_2\text{O} \longrightarrow 2\text{NaOH} + \text{H}_2$ $\text{Na}^0 \text{ to } \text{Na}^{+}$; oxidized / R.A. $\text{H}^{+} \text{ to } \text{H}^0$; reduced / O.A.

11. Give the oxidation number of each kind of atom or ion.

- | | | | | | | |
|--|------------|--------------------------|---------------------------|---------------------------|---------------------------------------|--|
| a. sulfate
SO_4^{2-}
2- | b. Sn
0 | c. S^{2-}
2- | d. Fe^{3+}
3+ | e. Sn^{4+}
4+ | f. nitrate
NO_3^{-}
1- | g. ammonium
NH_4^{+}
1+ |
|--|------------|--------------------------|---------------------------|---------------------------|---------------------------------------|--|

12. Calculate the oxidation number of chromium in each of the following. (work on next page)

- | | | | | |
|----------------------------------|---|--------------------------|-------------------|---------------------|
| a. Cr_2O_3
3+ | b. $\text{Na}_2\text{Cr}_2\text{O}_7$
6+ | c. CrSO_4
2+ | d. chromate
6+ | e. dichromate
6+ |
|----------------------------------|---|--------------------------|-------------------|---------------------|

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)

- $\text{C} + \text{H}_2\text{SO}_4 \longrightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$ $\text{C}^0 \text{ to } \text{C}^{4+}$; oxidized / R.A. $\text{S}^{6+} \text{ to } \text{S}^{4+}$; reduced / O.A.
- $\text{HNO}_3 + \text{HI} \longrightarrow \text{NO} + \text{I}_2 + \text{H}_2\text{O}$ $\text{N}^{5+} \text{ to } \text{N}^{2+}$; reduced / O.A. $\text{I}^{-} \text{ to } \text{I}^0$; oxidized / R.A.
- $\text{KMnO}_4 + \text{HCl} \longrightarrow \text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O} + \text{KCl}$ $\text{Mn}^{7+} \text{ to } \text{Mn}^{2+}$; reduced / O.A. $\text{Cl}^{-} \text{ to } \text{Cl}^0$; oxidized / R.A.
- $\text{Sb} + \text{HNO}_3 \longrightarrow \text{Sb}_2\text{O}_3 + \text{NO} + \text{H}_2\text{O}$ $\text{Sb}^0 \text{ to } \text{Sb}^{3+}$; oxidized / R.A. $\text{N}^{5+} \text{ to } \text{N}^{2+}$; reduced / O.A.

14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

Chapter 20 Redox Reactions Answers

John T. Moore, Richard H. Langley



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touché dans son amour propre il laisse l orgueil et la vanité le dominer et décide de la détromper en séduisant rosette une jeune paysanne sœur de lait de camille espérant ainsi rendre sa cousine jalouse lui donnant rendez vous afin qu elle assiste à la scène

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publiée pour la première fois en 1834 on ne badine pas avec l amour est une des œuvres les plus appréciées d alfred de musset notamment connue pour être le reflet de sa relation avec l auteure george sand pur mélange entre comédie et tragédie et entre réalité et fiction cette pièce a connu une première adaptation

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le drame des rêves de la jeunesse conclusion le drame de l orgueil camille et perdican en effet étaient promis l un à l autre depuis longtemps et la rencontre trop habilement combinée par le baron devait préluder à un mariage heureux mais chacun d eux repousse à son tour l heureux projet en se laissant aller à un mouvement de vanité

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celui ci arrive à ses fins mais lors de leur premier baiser rosette amoureuse de perdican les aperçoit et pousse un cri strident rosette se tue et camille mise au courant du stratagème de perdican le quitte le titre sonne comme un avertissement on

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alfred de musset 1861 la pièce se déroule au château du baron camille la nièce du baron sort du couvent elle a 18 ans elle retrouve le fils de son oncle perdican qui a 21 ans et vient d obtenir un doctorat ils ne s étaient pas vus depuis dix ans le baron souhaite les marier perdican et camille s aiment depuis toujours