

Chapter 10: Thermodynamics

The First Law of Thermodynamics

Q10.1. An uncharged, adiabatic piston (with 1 mol of gas) is compressed isothermally (200 K) and slowly (quasi-statically) to one-half its volume. What is the increase in the internal energy of the system?

Indicate the sign (positive or negative) before the correct numerical answer. Apply the first

$$\Delta U = \Delta Q + \Delta W \quad \Delta Q = \text{heat in} \quad \Delta W = \text{work in}$$

$$\Delta U = \Delta Q + \Delta W = \text{heat in} + \text{work in} = \boxed{\Delta U = \text{heat in}}$$

Q10.2. Assume that this process is quasi-static (the piston is moved by 100 N while 200 J of work is done on it) as you indicate in the table. What is the heat ΔQ (in Joules) required by the process?

$$\Delta U = \Delta Q + \Delta W \quad \Delta W = \text{work in} \quad \Delta Q = \text{heat in} + \Delta W$$

$$\Delta Q = \Delta U - \Delta W = 1,000 \text{ J} - 1,200 \text{ J} = \boxed{\Delta Q = \text{heat in} - \Delta W}$$

Q10.3. In thermodynamic processes, the internal energy of the system corresponding to 100 J of heat added to the reservoir by the gas (1000 J of heat is absorbed)

$$\Delta U = \Delta Q + \Delta W \quad \Delta Q = \text{heat in} \quad \Delta W = \text{work in} + \Delta U$$

$$\Delta W = \Delta Q - \Delta U = \text{heat in} - \text{heat in} = \boxed{\Delta W = \text{heat in}}$$

Q10.4. Suppose that 100 J of heat is added to a system, which has a corresponding process 100 J of work done on the system. What is the change in internal energy of the system? Indicate the

correct [] internal $\Delta U = 0$, where $\Delta Q = \Delta U + \Delta W = 0$ and $\Delta W = 0$ and $\Delta U = 0$

$$\Delta U = \Delta Q + \Delta W = \text{heat in} + \Delta W = 1,000 \text{ J} + 1,000 \text{ J} = \boxed{\Delta U = \text{heat in}}$$

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David Halliday, Robert Resnick, Jearl Walker



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