

## Chapter 20 Worksheets: Redox ANSWERS

1. Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent, also.

- $2\text{Sr} + \text{O}_2 \longrightarrow 2\text{SrO}$   $\text{Sr}^0 \text{ to } \text{Sr}^{2+}$ ; oxidized/reducing agent  $\text{O}^0 \text{ to } \text{O}^{2-}$ ; reduced/ox. ag.
- $2\text{Li} + \text{S} \longrightarrow \text{Li}_2\text{S}$   $\text{Li}^0 \text{ to } \text{Li}^{+}$ ; oxidized/ox. ag.  $\text{S}^0 \text{ to } \text{S}^{2-}$ ; reduced/ox. ag.
- $2\text{Cu} + \text{Br}_2 \longrightarrow 2\text{CuBr}$   $\text{Cu}^0 \text{ to } \text{Cu}^{+}$ ; oxidized/ox. ag.  $\text{Br}^0 \text{ to } \text{Br}^{-}$ ; reduced/ox. ag.
- $3\text{Mg} + \text{N}_2 \longrightarrow \text{Mg}_3\text{N}_2$   $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$ ; oxidized/ox. ag.  $\text{N}^0 \text{ to } \text{N}^{3-}$ ; reduced/ox. ag.
- $4\text{Fe} + 3\text{O}_2 \longrightarrow 2\text{Fe}_2\text{O}_3$   $\text{Fe}^0 \text{ to } \text{Fe}^{3+}$ ; oxidized/ox. ag.  $\text{O}^0 \text{ to } \text{O}^{2-}$ ; reduced/ox. ag.
- $\text{Cl}_2 + 2\text{NaBr} \longrightarrow 2\text{NaCl} + \text{Br}_2$   $\text{Cl}^0 \text{ to } \text{Cl}^{-}$ ; reduced/ox. ag.  $\text{Br}^{-} \text{ to } \text{Br}^0$ ; oxidized/ox. ag.
- $\text{Si} + 2\text{F}_2 \longrightarrow \text{SiF}_4$   $\text{Si}^0 \text{ to } \text{Si}^{4+}$ ; oxidized/ox. ag.  $\text{F}^0 \text{ to } \text{F}^{-}$ ; reduced/ox. ag.
- $\text{Mg} + 2\text{HCl} \longrightarrow \text{MgCl}_2 + \text{H}_2$   $\text{Mg}^0 \text{ to } \text{Mg}^{2+}$ ; oxidized/ox. ag.  $\text{H}^{+} \text{ to } \text{H}^0$ ; reduced/a.
- $2\text{Na} + 2\text{H}_2\text{O} \longrightarrow 2\text{NaOH} + \text{H}_2$   $\text{Na}^0 \text{ to } \text{Na}^{+}$ ; oxidized/a.  $\text{H}^{+} \text{ to } \text{H}^0$ ; reduced/a.

11. Give the oxidation number of each kind of atom or ion.

- |            |       |                    |                     |                     |            |             |
|------------|-------|--------------------|---------------------|---------------------|------------|-------------|
| a. sulfate | b. Sn | c. $\text{S}^{2-}$ | d. $\text{Fe}^{3+}$ | e. $\text{Sn}^{4+}$ | f. nitrate | g. ammonium |
| 2-         | 0     | 2-                 | 3+                  | 4+                  | 1-         | 1+          |

12. Calculate the oxidation number of chromium in each of the following.

- |                            |                                       |                    |             |               |
|----------------------------|---------------------------------------|--------------------|-------------|---------------|
| a. $\text{Cr}_2\text{O}_3$ | b. $\text{Na}_2\text{Cr}_2\text{O}_7$ | c. $\text{CrSO}_4$ | d. chromate | e. dichromate |
| 3+                         | 6+                                    | 2+                 | 7+          | 6+            |

13. Use the changes in oxidation numbers to determine which elements are oxidized and which are reduced in these reactions. (Note: it is not necessary to use balanced equations)

- $\text{C} + \text{H}_2\text{SO}_4 \longrightarrow \text{CO}_2 + \text{SO}_2 + \text{H}_2\text{O}$   $\text{C}^0 \text{ to } \text{C}^{4+}$ ; oxidized  $\text{S}^{6+} \text{ to } \text{S}^{4+}$ ; reduced
- $\text{HNO}_3 + \text{HI} \longrightarrow \text{NO} + \text{I}_2 + \text{H}_2\text{O}$   $\text{N}^{5+} \text{ to } \text{N}^{2+}$ ; reduced  $\text{I}^{-} \text{ to } \text{I}^0$ ; oxidized
- $\text{KMnO}_4 + \text{HCl} \longrightarrow \text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O} + \text{KCl}$   $\text{Mn}^{7+} \text{ to } \text{Mn}^{2+}$ ; reduced  $\text{Cl}^{-} \text{ to } \text{Cl}^0$ ; oxidized
- $\text{Sn} + \text{HNO}_3 \longrightarrow \text{SnO}_2 + \text{NO} + \text{H}_2\text{O}$   $\text{Sn}^0 \text{ to } \text{Sn}^{4+}$ ; oxidized  $\text{N}^{5+} \text{ to } \text{N}^{2+}$ ; red.

14. For each reaction in problem 13, identify the oxidizing agent and reducing agent.

- |                               |                          |
|-------------------------------|--------------------------|
| a. oxidizing agent: sulfur    | reducing agent: carbon   |
| b. oxidizing agent: nitrogen  | reducing agent: iodine   |
| c. oxidizing agent: manganese | reducing agent: chlorine |
| d. oxidizing agent: nitrogen  | reducing agent: antimony |

## Chapter 20 Study Guide Redox Reactions Answers

**Holder**



## **Chapter 20 Study Guide Redox Reactions Answers:**

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## **Chapter 20 Study Guide Redox Reactions Answers** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the energy of words has are more evident than ever. They have the ability to inspire, provoke, and ignite change. Such is the essence of the book **Chapter 20 Study Guide Redox Reactions Answers**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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