

CHAPTER 25

NUCLEAR CHEMISTRY

PRACTICE EXAMPLES

1A (E) A β^- has a mass number of zero and an “atomic number” of -1 . Emission of this electron has the effect of transforming a neutron into a proton. ${}_{94}^{241}\text{Pu} \rightarrow {}_{95}^{241}\text{Am} + {}_{-1}^0\beta$

1B (E) ${}^{58}_{28}\text{Ni}$ has a mass number of 58 and an atomic number of 28. A positron has a mass number of 0 and an effective atomic number of $+1$. Emission of a positron has the seeming effect of transforming a proton into a neutron. The parent nuclide must be copper-58.



2A (E) The sum of the mass numbers ($139 + 12 = ? + 147$) tells us that the other product species has $A = 4$. The atomic number of La is 57, that of C is 6, and that of Eu is 63. The atomic number sum ($57 + 6 = ? + 63$) indicates that the atomic number of this product species is zero. Therefore, four neutrons must have been emitted.



2B (E) An alpha particle is ${}_2^4\text{He}$ and a positron is ${}_{+1}^0\beta$. We note that the total mass number in the first equation is 125; the mass number of the additional product is 1.

The total atomic number is 53; the atomic number of the additional product is 0; it is a neutron. ${}^{125}_{53}\text{Sb} + {}_2^4\text{He} \rightarrow {}^{124}_{52}\text{I} + {}_0^1n$

In the second equation, the positron has a mass number of 0, meaning that the mass number of the product is 124. Because the atomic number of the positron is $+1$, that of the product is 52; it is ${}^{124}_{52}\text{Te}$.



3A (M) (a) The decay constant is found from the 8.040-day half-life.

$$\lambda = \frac{0.693}{8.040 \text{ d}} = 0.0862 \text{ d}^{-1} \times \frac{1 \text{ d}}{24 \text{ h}} \times \frac{1 \text{ h}}{60 \text{ min}} \times \frac{1 \text{ min}}{60 \text{ s}} = 9.98 \times 10^{-7} \text{ s}^{-1}$$

(b) The number of ${}^{131}\text{I}$ atoms is used to find the activity.

$$\begin{aligned} \text{no. } {}^{131}\text{I} \text{ atoms} &= 2.05 \text{ mg} \times \frac{1 \text{ g}}{1000 \text{ mg}} \times \frac{1 \text{ mol } {}^{131}\text{I}}{131 \text{ g } {}^{131}\text{I}} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mol } {}^{131}\text{I}} \\ &= 9.42 \times 10^{18} \text{ atoms } {}^{131}\text{I} \end{aligned}$$

$$\text{activity} = \lambda N = 9.98 \times 10^{-7} \text{ s}^{-1} \times 9.42 \times 10^{18} \text{ atoms} = 9.40 \times 10^{12} \text{ disintegrations / second}$$

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United States. Office of Saline Water



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