## Acids and Bases Problems

Rank the following solutions in order of increasing acidity (from lowest to highest):

Solution A with pH = 6.50 Solution B with  $[H_3O^*] = 3.5x10^{-5} M$  Solution D with pH = 5.85

Solution C with [OH-] = 6.0x10-10 M

\_\_ < \_\_ < \_\_ < \_\_

Complete the following table.

Acid	Conjugate Base
HS-	
	CIO <sub>4</sub> -
H <sub>3</sub> PO <sub>4</sub>	
	HCIO <sub>3</sub> -

3 Write the acid-base reactions with water for each of the following acids. Label the acid, base, conjugate acid, and conjugate base.

i) [Co(NH<sub>3</sub>)<sub>5</sub>(OH<sub>3</sub>)]3\*

HSO,

iii) CH<sub>2</sub>OH

W) H,PO,

# **Chapter 2acids Bases Practice Problems Answers**

**Steve Russo, Mike Silver** 

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