

**Ch. 9 Gases: Their Properties and Behavior****9.1 Gases and Gas Pressure**

gases expand to fill a container.

if mix of several compounds or elemental gases – homogeneous mixture (solution).

atoms or molecules – very far apart from each other mostly empty space.

density of water vapor: 0.803 g/L

density of water: 1000 g/L liquid more than 1000 times as dense as the gas...

$$\text{pressure (P)} = \frac{\text{Force (F)}}{\text{Area (A)}}$$

Pressure at sea level called 1 atmosphere (atm)

1 atmosphere of pressure pushes a column of mercury up 760 mm (or water 10.6 meters – 34.8 feet)

**9.2 Gas Laws****Boyle's Law**

If pressure increases, the volume of a gas decreases.

$$V \propto \frac{1}{P} \quad \text{or} \quad V = \frac{1}{P} \times \text{constant} \quad \text{or} \quad PV = \text{constant}$$

**Charles Law:**

If T increases, average velocity of atoms/molecules of gas increase – greater force from collisions – If the pressure stays constant, the volume will increase.

$$V \propto T \quad \text{or} \quad V = T \times \text{constant} \quad \text{or} \quad \frac{V}{T} = \text{constant}$$

**NOTE:** For Charles Law, T must be in an absolute temperature scale. For SI units: Kelvin

**Avogadro's law:**

If the number of gas atoms/molecules increased, volume increases (if P constant).

$$V \propto n \quad \text{or} \quad V = n \times \text{constant} \quad \text{or} \quad \frac{V}{n} = \text{constant}$$

where n = moles of gas

**9.3 The Ideal Gas Law**

Combining all 3 relations:

$$V \propto n$$

$$V \propto T$$

# Chapter 1review Gases Section Worksheet Answers

**Camilla Rothe**



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