

## Chapter 1: Solutions to Review Problems

### Section 1.1

- 1) If the number of moles of gas and the temperature remain constant:  $P_1V_1 = P_2V_2$

$$\text{Substitute: } 100 \text{ kPa} \times 2.00 \text{ L} = P_2 \times 3.00 \text{ L} \therefore P_2 = 66.7 \text{ kPa}$$

- 2) If the number of moles of gas and the pressure remain constant:  $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

$$\text{Substitute: } \frac{1.00 \text{ L}}{200 \text{ K}} = \frac{V_2}{400 \text{ K}} \quad \therefore V_2 = 2.00 \text{ L}$$

- 3) If the number of moles of gas and the volume remain constant:  $\frac{P_1}{T_1} = \frac{P_2}{T_2}$

$$\text{Substitute: } \frac{100 \text{ kPa}}{298 \text{ K}} = \frac{P_2}{398 \text{ K}} \quad \therefore P_2 = 133.6 \text{ kPa}$$

- 4) Xe has molar mass  $131.3 \text{ g mol}^{-1}$ , Ar has molar mass  $39.95 \text{ g mol}^{-1}$ , Ne has molar mass  $20.18 \text{ g mol}^{-1}$

If equal masses of each gas are used, the sample of the gas with the smallest molar mass will contain the greatest number of moles.

Therefore, the sample of Ne will contain the greatest number of moles.

Pressure is directly proportional to the number of moles of gas; so the pressure in the Ne flask is the greatest. (Answer a)

- 5)  $\text{O}_2$  has molar mass  $32.00 \text{ g mol}^{-1}$ ,  $\text{N}_2$  has molar mass  $28.02 \text{ g mol}^{-1}$

If equal masses of each gas are used, the sample of the gas with the smaller molar mass will contain the greater number of moles.

Therefore, the sample of  $\text{N}_2$  will contain the greater number of moles, and therefore molecules.

Pressure is directly proportional to the number of moles (or molecules) of gas; so the pressure in the  $\text{N}_2$  flask is the greater than the pressure in the  $\text{O}_2$  flask. (Answer c)

# Chapter 1 solutions Test Answers

**Lial, Miller**



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