



1. List the five assumptions of the Kinetic Molecular Theory of gases.

- Gases have teeny, tiny particles far, far apart.
- The particles are in constant, rapid, random, straight-line motion.
- Collisions between particles are perfectly elastic.
- No forces of attraction or repulsion act between the particles.
- $KE = \frac{1}{2} m u^2$

2. Write the formula for Boyle's law. $P_1V_1 = P_2V_2$

3. If the volume of a gas at constant temperature and at 735 torr changes from 75.0 mL to 90.0 mL, what is the new pressure?

$$P_2 = (P_1V_1 / V_2) = (735 \text{ torr} * 75.0 \text{ mL} / 90.0 \text{ mL}) = \underline{613 \text{ torr}}$$

4. Write the formula for Charles' law. $V_1 / T_1 = V_2 / T_2$

5. If the temperature of 15.0 L of a gas at constant pressure changes from -30.0°C to 25.0°C , what is the new volume?

$$V_2 = (T_2V_1 / T_1) = (298 \text{ K} * 15.0 \text{ L} / 243 \text{ K}) = \underline{18.4 \text{ L}}$$

6. Write the formula for Gay-Lussac's law. $P_1 / T_1 = P_2 / T_2$

7. If the temperature of a gas at 1.00 atm and contained in a fixed volume changes from 20.0°C to -15.0°C , find the new pressure.

$$P_2 = (T_2P_1 / T_1) = (258 \text{ K} * 1.00 \text{ atm} / 293 \text{ K}) = \underline{0.881 \text{ atm}}$$

Chapter 1the Gas Laws Answer Key

Wolfgang Guggemos

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