

## Chapter 20- Chemical Bonds

- I. Stability of Bonding
  - a. Some properties combine chemically and no longer have the same properties they did before forming a compound.
  - b. A(n) chemical formula is composed of symbols and subscripts indicating the number of atoms of an element in a compound.
  - c. Atoms form compounds when the compound is more stable than the separate atoms.
    - i. Noble gases are more chemically stable than other elements because they have a complete outer energy level.
    - ii. Atoms can lose, gain, or share electrons to get a stable outer energy level.
    - iii. A(n) chemical bond is the force that holds atoms together in a compound.
- II. Types of Bonds
  - a. A(n) ion is charged particle because it has more or fewer electrons than protons.
    - i. When an atom loses an electron, it becomes a positively charged ion; a superscript indicates the charge.
    - ii. When an atom gains an electron, it becomes a negatively charged ion.
  - b. An ionic compound is held together by the ionic bond- the force of attraction between opposite charges of the ions.
    - i. The result of this bond is a neutral compound.
    - ii. The sum of the charges on the ions is zero.
  - c. Molecules are neutral particles formed as a result of sharing electrons.
    - i. A covalent bond is the force of attraction between atoms sharing electrons.
    - ii. Atoms can form double or triple bonds depending on whether they share two or three pairs of electrons.
    - iii. Electrons shared in a molecule are held more closely to the atoms with the larger nucleus.
    - iv. A(n) polar molecule has one end that is slightly negative and one end that is slightly positive although the overall molecule is neutral.
    - v. In a(n) nonpolar molecule electrons are shared equally.
- III. Writing Formulas and Naming Compounds
  - a. Chemists use symbols from the periodic table to write formulas for compounds.
  - b. Binary Compounds- composed of two elements
    - i. Oxidation number- how many electrons an atoms has gained, lost, or shared to become stable
    - ii. Use oxidation numbers and their least common multiples to write formulas.
      - 1. When writing formulas, remember that the compound is neutral.
      - 2. A formula must have the correct number of positive and negative ions so the charges balance.
    - iii. Use the name of the first element, the root name of the second element, and the suffix -*ide* to write the name of the binary ionic compound.
  - c. Polyatomic ion- positively or negatively charged, covalently bonded groups of atoms.
    - i. The compound contains three or more elements.
    - ii. To write names, write the name of the positive ion first; then write the name of the negative ion.
    - iii. To write formulas, use the oxidation numbers, their least common multiple, and put parentheses around the polyatomic ion before adding a subscript.
  - d. Hydrate- compound with water chemically attached to its ions
  - e. Name binary covalent compounds by using prefixes to indicate how many atoms of each element are in the compound.

# Chapter 20 Chemical Bonds Directed Reading Answer Key

**Neil A. Campbell, Jane B. Reece**



## Chapter 20 Chemical Bonds Directed Reading Answer Key:

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