

Chapter 21 -- Nuclear Chemistry

21.1 Radioactivity

- When nuclei change spontaneously, emitting energy, they are said to be **radioactive**.
- **Nuclear chemistry** is the study of nuclear reactions and their uses.
- **Isotopes** have the same number of p⁺ but different numbers of n⁰.
 - Different isotopes of the same element are distinguished by their mass numbers.
 - Different isotopes have different natural abundances.
- A **radionuclide** is a radioactive nucleus.
 - Atoms containing these nuclei are called **radioisotopes**.

Nuclear Equations

- Most nuclei are stable.
 - Radionuclides are unstable and spontaneously emit particles and/or electromagnetic radiation.
 - Example: Uranium-238 is radioactive.
 - It emits **alpha (α) particles** (helium-4 nuclei)
- When a nucleus spontaneously decomposes in this manner, we say it has decayed (*radioactive decay*).
- In nuclear equations, the total number of nucleons is conserved.
 - We can represent the uranium-238 decay by the following nuclear equation:



- The total number of protons and neutrons before a nuclear reaction must be the same as the total number of nucleons after the reaction.

Types of Radioactive Decay

- There are three types of radiation which we will consider:
 - α-Radiation is the loss of ${}_2^4\text{He}$ (**alpha particles**) from the nucleus.
 - β-Radiation is the loss of an electron from the nucleus. These high-speed electrons are called **beta particles**.
 - γ-Radiation is the loss of high-energy photons from the nucleus.
- Nucleons can undergo two other types of decay:
 - **positron emission**
 - A positron is a particle with the same mass as an electron but with the opposite sign.
 - **electron capture**
 - The nucleus captures an electron from the electron cloud surrounding the nucleus.
- Representations:
 - In nuclear chemistry, to ensure the conservation of nucleons we write all particles with their atomic and mass numbers: ${}_2^4\text{He}$ and ${}_{-1}^0\text{e}$ represent α- and β- radiation.
- Nucleons can undergo decay:
 - (β-emission) ${}_0^1\text{n} \rightarrow {}_1^1\text{p} + {}_{-1}^0\text{e}$
 - (positron annihilation) ${}_1^0\text{e} + {}_{-1}^0\text{e} \rightarrow 2{}_0^0\gamma$
 - (positron or β⁺-emission) ${}_1^1\text{p} \rightarrow {}_0^1\text{n} + {}_1^0\text{e}$
 - (electron capture) ${}_1^1\text{p} + {}_{-1}^0\text{e} \rightarrow {}_0^1\text{n}$

Chapter 21 Nuclear Chemistry Test Answer Key

Charles H. Corwin



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An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

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SAT II Claire Aldridge, Karla Lee, Kaplan, Inc, 2000 Expert review prepares test takers for all the chemistry content they need to know with three full length practice tests and detailed explanations Charts *Student's Guide, Chemistry, the Central Science* James C. Hill, 1991 **General Chemistry** Darrell D. Ebbing, 1999 *Kaplan SAT II Chemistry 2002-2003* Claire Aldridge, Karl Lee, Kaplan, Inc, 2002 Hundreds of practice questions including 3 full length practice tests and quizzes in every chapter Proven strategies to help students score higher *Student's Guide to Brown and LeMay Chemistry* James C. Hill, 1988 Kaplan SAT II: Chemistry 2004-2005 Kaplan, 2004-03-02 At head of title Kaplan **Chemistry - The Central Science** James C. Hill, Bruce Edward Bursten, 2006 Chemistry The Central Science is the most trusted book on the market its scientific accuracy clarity innovative pedagogy functional problem solving and visuals set this book apart Brown LeMay and Bursten teach students the concepts and skills they need without overcomplicating the subject A comprehensive media package that works in tandem with the text helps students practice and learn while providing instructors the tools they need

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