Review Definitions

Chemical Energy is energy stored in the bonds of atoms and molecules.

Nuclear Energy is due to changes in the nucleus of atoms changing them into different atoms. Nuclear energy is enormous in comparison.

#### The Nucleus

Mass rumber (number of protons plus neutrons)

Abomic number (number of protons or electrons)

Nucleus is composed of the two nucleons, protons and neutrons.

The number of protons is the atomic number.

The number of protons and neutrons together is the mass number.

**Isotopes:** the same element have the different mass, due to different numbers of neutrons in those atoms.

Radioactivity: some nuclides of an elements are unstable, or radioactive, which can be called radionuclides (放射性核素).

Property	Type of Radiation		
	- 100	All I	7
Charge	2+	1-	0
Mass	$6.64 \times 10^{-24} \mathrm{g}$	$9.11 \times 10^{-26} g$	0
Relative penetrating power	1	100	10,000
Nature of radiation	Diffe muchei	Ethyot rouses	High energy photons

Types of Radioactive Decay:

Alpha decay is the loss of an α-particle.

Beta decay is the loss of a β-particle (a high-speed electron emitted by the nucleus): <sup>0</sup><sub>-1</sub>β or <sup>0</sup><sub>-1</sub>e

 Gamma emission is the loss of a y-ray, which is high-energy radiation that almost always accompanies the loss of a nuclear particle: <sup>0</sup>Y

# **Chapter 21 Nuclear Chemistry Worksheets**

**M** Walker

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  - Chapter 2: Essential Elements of Chapter 21 Nuclear Chemistry Worksheets
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  - Chapter 4: Chapter 21 Nuclear Chemistry Worksheets in Specific Contexts
  - ∘ Chapter 5: Conclusion
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- 6. In chapter 5, the author will draw a conclusion about Chapter 21 Nuclear Chemistry Worksheets. This chapter will summarize the key points that have been discussed throughout the book.

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