

DAY 4
BIO 112 NOTE-TAKING AND STUDY GUIDE
WATER AND CARBON
CHAPTERS 2.3 AND CHAPTER 2.4

Section 2.3 WATER

Water is a polar molecule.

What's the difference between polar and nonpolar molecules?

Polar Molecules: Have a partial positive charge on one end and a partial negative charge on the other, due to uneven distribution of electrons. This is often due to the presence of electronegative atoms like oxygen or nitrogen. For example, water (H_2O) is polar because oxygen is more electronegative than hydrogen, creating a dipole moment.

Nonpolar Molecules: Have an even distribution of electrons, resulting in no partial charges. They do not have dipole moments. Examples include gases like nitrogen (N_2) and carbon dioxide (CO_2).

What do hydrophilic and hydrophobic mean?

Hydrophilic: Water-loving; these molecules are polar or charged and can form hydrogen bonds or interact with water. Examples include salts and sugars.

Hydrophobic: Water-fearing; these molecules are nonpolar and do not interact well with water. Examples include oils and fats.

What is the hydrophobic effect?

The hydrophobic effect refers to the tendency of nonpolar substances to collective in aqueous solutions and exclude water molecules. This phenomenon drives the formation of structures like lipid bilayers in cell membranes.

A hydrogen bond is an interaction between a hydrogen atom and an electronegative atom.

Define hydrogen bond

Study Fig. 2.10

A hydrogen bond is a type of weak, non-covalent interaction that occurs between a hydrogen atom covalently bonded to a more electronegative atom (like oxygen or nitrogen) and another electronegative atom. This interaction is crucial in the structure and function of biological molecules.

Hydrogen bonds give water many unusual properties.

What are some of the unusual properties of water?

High Specific Heat: Water can absorb a lot of heat before its temperature changes significantly.

High Heat of Vaporization: It requires a lot of energy to convert water from liquid to vapor.

High Surface Tension: Water molecules stick together more than to air, allowing it to form droplets and support small objects.

Expands Upon Freezing: Ice is less dense than liquid water, which is why ice floats.

Solvent Abilities: Water is a versatile solvent due to its polarity, dissolving many ionic and polar substances.

pH is a measure of the concentration of protons in solution.

What is the formula for pH?

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