

Chemistry Unit 1 ISA 2009 – 1 hour

SECTION A

These questions are about the task, the composition of antacid tablets. You should use the Task Sheet and your own Candidate Results Sheet to answer them.

1 Record the average titre from your Candidate Results Sheet. (1 mark)

2 The equation for the reaction of sodium hydrogencarbonate and hydrochloric acid is shown below.



The concentration of the hydrochloric acid used was $0.100 \text{ mol dm}^{-3}$. Use your answer from Q1 to calculate the concentration, in mol dm^{-3} , of sodium hydrogencarbonate in the antacid solution. (2 marks)

3 Use data from the Periodic Table to calculate the relative formula mass of sodium hydrogencarbonate. (1 mark)

4 Use your answers from question 2 and question 3 to determine the concentration, in g dm^{-3} , of sodium hydrogencarbonate in the antacid solution. (1 mark)

5 The antacid solution was prepared by dissolving 20.0 g of the tablets in water and making up to 1.00 dm^3 of solution. Use your answer from question 4 to calculate the percentage by mass of sodium hydrogencarbonate in the tablets. (1 mark)

6 The maximum total error in using the burette is $\pm 0.15 \text{ cm}^3$. This error takes into account multiple measurements. Use the average titre from question 1 to calculate the percentage error in using the burette. (1 mark)

7 Sucrose is used in antacid tablets to improve their taste. Suggest one reason why the presence of sucrose in the antacid tablet does not affect the titration with hydrochloric acid. (1 mark)

SECTION B

INTRODUCTION

In Section A you determined the percentage by mass of sodium hydrogencarbonate in antacid tablets. Some antacid tablets contain sodium hydrogencarbonate, sucrose and citric acid.

8 Analysis of a pure sample of citric acid showed that it contained 37.50% of carbon and 4.17% of hydrogen by mass, the remainder being oxygen. Use these data to show that the empirical formula of the acid is $\text{C}_6\text{H}_8\text{O}_7$. (3 marks)

9 When the antacid tablet is added to water, sodium hydrogencarbonate and citric acid react together to form a gas. Identify this gas. (1 mark)

10 A weighed portion of this antacid was added to water. The gas formed was collected and its volume measured.

(a) Draw a diagram to show how this experiment could have been carried out to collect and measure the volume of the gas. (2 marks)

(b) The experiment was repeated with further weighed portions of the same antacid. The results are shown.

Experiment	1	2	3	4	5
Mass of antacid	2.60	1.17	0.88	2.31	1.80
Volume of gas collected / cm^3	168	86	57	149	116

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