

Chapter 1 – Kinetics - Practice Test #1

Name - _____

The answers to this test are posted online

1. Consider the following reaction mechanism:



The chemical species MX is a(n)

- A. catalyst
B. inhibitor
C. final product
D. reaction intermediate
2. Consider the following reaction: $2\text{N}_2\text{O}_{5(g)} \rightarrow 4\text{NO}_{2(g)} + \text{O}_{2(g)}$
At a certain temperature the rate of decomposition of $2\text{N}_2\text{O}_{5(g)}$ is 2.5×10^{-6} mol/s. The rate of formation of NO_2 is
- A. 1.0×10^{-5} mol/s
B. 1.3×10^{-6} mol/s
C. 2.5×10^{-6} mol/s
D. 5.0×10^{-6} mol/s

3. Which of the following factors affect the rates of both homogeneous and heterogeneous reactions.

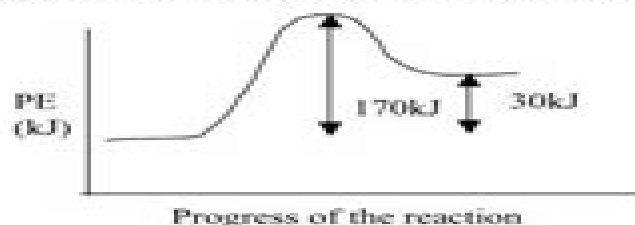
- I nature of the reactants
II presence of a catalyst
III temperature of system
IV concentration of reactants

- A. I and IV only
B. II and III only
C. II, III, and IV only
D. I, II, III, and IV
4. Which of the following equations represents an endothermic reaction?

- A. $\text{N}_2\text{O}_{4(g)} + 59 \text{ kJ} \rightarrow 2\text{NO}_{2(g)}$
B. $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(l)} + 572 \text{ kJ}$
C. $2\text{BrCl}_{(g)} - 29.3 \text{ kJ} \rightarrow \text{Br}_{2(g)} + \text{Cl}_{2(g)}$
D. $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(l)} \quad \Delta H = -572 \text{ kJ}$

5. Consider the potential energy diagram. The activation energy for the reverse reaction is

- A. 30 kJ
B. 140 kJ



- C. 170 kJ
D. 200 kJ

Chapter 1 chemical Kinetics Test

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